Level - I (Based on NCERT)

 Which is not an example of exothermic reaction: (a) N₂ + O₂ → 2NO - Heat (b) N₂ + 3H₂ →2NH₃ + Heat (c) H₂ + I₃ →2HI + Heat (d) All of these 2. Which is an example of endothermic reaction: (a) N₂ + O₂ →2NO - Heat (b) CaCO₃ → CaO + CO₂ − Heat (c) 2KClO₃ → 2kCl + 3O₂ − Heat (d) All of these Which is an example of combination reaction: (a) N₂ + Cl₂ → 2HCl (b) C + O₂ → CO₂ (d) All of these (c) $N_2 + 3 H_2 \rightarrow 2NH_3$ 4. Which is an example of decomposition reaction? (a) 2Pb(NO₃)₂ → 2PbO + 4NO₂ + O₂ (b) 2FeSO₄ → Fe₂O₄ + SO₂ + SO₃ (c) 2H2O Electricity 2H2+ O2 (d) All of these 5. Which is an example of thermal decomposition? (b) 2Al₂O₃ Electricity 4Al + 3O₃ (a) 2NaCl Electricity 2Na + Cla (c) CaCO₃ → CaO + CO₂ (d) NaOH + HCl →NaCl + H₂O 6. Which is an example of electrolysis decomposition? (a) 2KClO₃ → 2KCl + 3O₂ (b) BaCl₂ + Na₂SO₄ → BaSO₄ + 2NaCl (d) 2Al₂O₂(I) Electricity 4Al + 3O. (c) 2AgCl hv2Ag + Cl2 7. Which is an example of photo chemical reactions? (a) H₂ + C12 Ker 2 HC1 (b) 2AgBr W 2Ag + Br2 (c) 6CO2 + 6H2O W C6H12O6 (d) All of these 8. Which compound is used in photography? (b) AgCI (c) Na₂CO₂ (d) (a) and (b) 9. Which is an example of double displacement reaction? (a) BaCl₂ + Na₂SO₄ → BaSO₄ + 2NaCl (b) AgNO₃ + NaCl →AgCl + NaNO₃ (c) NaOH + HCl →NaCl + H₂O (d) All of these 10. Fe₂O₃ + 2Al →Al₂O₃ + 2Fe is an example of (a) Combination reaction (b) Decomposition reaction (c) Double displacement reaction (d) Displacement reaction 11. Which is an example of double displacement reaction? (a) Mg + 2HCl → MgCl₂ + H₂ (b) ZnCO₃ → ZnO + CO₂ (c) H₂ + Cl₂ →2HCl (d) BaCl₂ + K₂SO₄ → BaSO₄2KCl 12. H₂S + Cl₂ →2HCl + S is this reaction 'Cl₂' act as: (a) Oxidising agent (b) Reducing agent (d) None of these (c) Both (a) & (b) Which statement is correct regarding the reaction 2PbO + C → C₂Pb + CO₂ (a) Lead is getting reduced (b) Carbon dioxide is getting oxidised (c) Lead oxide is getting reduced (d) All of these 14. In the reaction MnO₂ + 4HCl → MnCl₂ + 2H2O + Cl₂, MnO₂ act as? (a) Reducing agent (b) Oxidising agent (c) Both (a) & (b) (d) None of these 15. Which is an example of redox reaction? (a) Zn + CuSO₄ → ZnSO₄ + Cu (b) H₂S + I2 → 2HCl + S (c) CuO + H₂ → Cu + H₂O (d) All of these 16. Which is a formula of rust? (a) Fe₂O₃ (b) Fe₂O (c) Fe₂O₃×H₂O (d) FeCO₁ 17. Rusting of Iron is an example: (a) Oxidation (b) Reduction (c) Redox (d) All of these 18. Antioxidant is/are: (d) None of these (a) Oxidising agent (b) Reducing agent (c) Both (a) & (b) 19. Rancidity can be prevented by: (a) Adding antioxidants (b) Storing food in air light

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Rachel Sandford

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web apr 24 2023 the easiest way to protect against uniform attack corrosion is usually to put a protective barrier between the metal and the corroding agents this can be a wide variety of things paint an oil sealant or an electrochemical solution like a galvanized zinc coating

17 5 corrosion and its prevention chemistry libretexts - Aug 15 2023

web aug 14 2020 in this section we describe some of the chemical and electrochemical processes responsible for corrosion we also examine the chemical basis for some common methods for preventing corrosion and treating corroded metals

corrosion is a redox process

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web structural and architectural design plays a major role to prevent corrosion and they shall know how to prevent corrosion the method of detailing the structural element shall be done in a manner that it does not encourage corrosion the following key aspect can be highlighted adequate access

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web jan 1 2022 introduction corrosion is a naturally occurring process which is defined as the degradation or deterioration of a substance and or its properties usually a metal over a period of time due to environmental exposure 1 this is an exergonic process as the metal tends toward the lowest possible energy state