Name: ______ Date: _____ Pd:____

MIXED Stoichiometry Practice Review Problems

STEP 1: The following reaction will be used for this entire worksheet, please complete the balancing before going on.

$$_1$$
Al₂(SO₄)₃ + $_3$ Ca(OH)₂ \rightarrow $_2$ Al(OH)₃ + $_3$ CaSO₄
342.17 g/mol 74.10 g/mol 78.01 g/mol 136.15 g/mol

STEP 2: Calculator the molar mass of each compound in the balanced reaction and write the values above the balanced reaction.

STEP 3: For aluminum sulfate, find the % of the metal in the compound. Please be sure to show your work. (Hint: W5.1 work)

Al is the only metal, S and O are nonmetals. There are no metalloids in the formula.

96 Al =
$$\frac{total\ mass\ Al}{molar\ mass\ Al_2(SO_4)_3}$$
 x 100 % = $\frac{53.96\ g}{342.17\ g}$ x 100 % = 15.77 %

STEP 4: Complete the following problems using a factor chart with units & chemicals. These problems could require ONE ratio, TWO ratios or even THREE ratios. Look for clues & use your mole map if necessary.

PROBLEMS

How many grams of calcium sulfate are produced from reacting 2.59 moles of aluminum sulfate?

2.59 mol Al₂(SO₄)₃ x
$$\frac{3 mol CaSO_4}{1 mol Al_2(SO_4)_3}$$
 x $\frac{136 15 g CaSO_4}{1 mol CaSO_4}$ = 1057.89 g CaSO₄

moles A problem B grams B

Mixed Stoichiometry Practice

RM Cervero

Mixed Stoichiometry Practice:

The Enigmatic Realm of Mixed Stoichiometry Practice: Unleashing the Language is Inner Magic

In a fast-paced digital era where connections and knowledge intertwine, the enigmatic realm of language reveals its inherent magic. Its capacity to stir emotions, ignite contemplation, and catalyze profound transformations is nothing lacking extraordinary. Within the captivating pages of **Mixed Stoichiometry Practice** a literary masterpiece penned with a renowned author, readers attempt a transformative journey, unlocking the secrets and untapped potential embedded within each word. In this evaluation, we shall explore the book is core themes, assess its distinct writing style, and delve into its lasting affect the hearts and minds of those that partake in its reading experience.

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