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## ACID-BASE EQUILIBRIA

### Arrhenius Acid:

Anything that produces hydrogen ions in aqueous solution

### Arrhenius Base:

Anything that produces hydroxide ions in aqueous solution

### Brønsted-Lowry Acid:

A species that donates a proton is an acid

### Brønsted-Lowry Base:

A species that accepts a proton is a base

**Amphiprotic Species:** a chemical species that can both donate and accept protons

### Lewis Acid:

Anything that accepts a pair of electrons

### Lewis Base:

Anything that donates a pair of electrons

### Water Equilibrium Constant ( $K_w$ ):

$K_w = [\text{H}_3\text{O}^+][\text{OH}^-] = 1.00 \times 10^{-14}$  (at 25°C)

### Strong Acid:

Acids that donate protons to water molecules quantitatively

### Strong Base:

A substance that generates hydroxide ions quantitatively in aqueous solution

### pH SCALES:

$\text{pH} = -\log[\text{H}_3\text{O}^+]$   $\text{pOH} = -\log[\text{OH}^-]$   $\text{p}K_a = -\log K_a$

$\text{p}K_b = -\log K_b$   $\text{pH} + \text{pOH} = 14.00$   $K_a K_b = K_w$   $\text{p}K_a + \text{p}K_b = 14.00$

### Weak Acid:

Acid that reaches equilibrium when only a small fraction of its molecules transfer protons to water

### Weak Base:

Generates hydroxide ions by accepting protons from water but reaches equilibrium when only a fraction of its molecules have done so

### APPLICATION OF Aqueous Equilibria:

#### BUFFER SOLUTION:

Contains both a weak acid and its conjugate base as major species in solution.

$$\text{pH} = \text{p}K_a + \log \left( \frac{[\text{A}^-]_{\text{total}}}{[\text{HA}]_{\text{total}}} \right)$$

#### BUFFER CAPACITY:

The amount of added  $\text{H}_3\text{O}^+$  or  $\text{OH}^-$  the buffer solution can tolerate without exceeding a specified pH range.

## CHEMICAL KINETICS

### REACTION MECHANISM:

The exact molecular pathway that starting materials follow on their way to becoming products

### RATE-DETERMINING STEP:

The slowest elementary step in a mechanism

### FIRST-ORDER RATE LAW:

Rate =  $k[A]$ , where A is a reactant in the overall reaction, this can be converted to

$$\ln \left( \frac{[A]_0}{[A]} \right) = kt$$

### HALF-LIFE ( $t_{1/2}$ ), 1<sup>st</sup> ORDER:

When half the original concentration has been consumed,

$$[A] = 0.5[A]_0 \quad t_{1/2} = \frac{\ln 2}{k}$$

### SECOND-ORDER RATE EXPRESSION:

Rate =  $k[A]^2$ , this can be converted to

$$\frac{1}{[A]} - \frac{1}{[A]_0} = kt$$

### BI-MOLECULAR ELEMENTARY REACTION:

$A + B \rightarrow \text{products}$  Elementary rate =  $k[A][B]$

### UNI-MOLECULAR ELEMENTARY REACTION:

$C \rightarrow \text{products}$  Elementary rate =  $k[C]$

### FIRST STEP IS RATE-DETERMINING:

The predicted rate law for the overall reaction is the rate expression for that first step

### ACTIVATION ENERGY ( $E_a$ ):

Energy barrier, the minimum energy that must be supplied before the reaction can occur

**Activated Complex:** the molecular arrangement at the point of highest energy along the energy level diagram

## CHEMICAL EQUILIBRIUM

### EQUILIBRIUM CONSTANT:

$$K_{eq} = \frac{[\text{D}]_{eq}^d [\text{E}]_{eq}^e}{[\text{A}]_{eq}^a [\text{B}]_{eq}^b}$$

$K_{eq}$  applied only at equilibrium,  $K_{eq}$  is independent of initial conditions,  $K_{eq}$  is related to the stoichiometry.

### Pure Liquid and Solid:

The concentrations of pure liquids or solids are always equal to their standard concentrations, therefore division by standard concentration results in a value of 1

### LARGE $K_{eq}$ :

Indicates that the reaction goes virtually to completion

### Le Chatelier's Principle:

When a change is imposed on a system at equilibrium, the system will react in the direction that reduces the amount of change.

### TEMPERATURE:

The only variable that causes a change in the value of  $K_{eq}$ , an increase in temperature always shifts the equilibrium position in the endothermic direction.

## THERMODYNAMICS

### CLOSED SYSTEM:

Exchanges energy but not matter with its surroundings

### ISOLATED SYSTEM:

Exchanges neither matter nor energy with the surroundings

### STATE VARIABLES:

Conditions that must be specified to establish the state of a system, pressure (P), volume (V), temperature (T), and amounts of substances (n)

### PHYSICAL CHANGE OF STATE:

Some of the state variables changes, but the chemical composition of the system stays the same

### CHEMICAL CHANGE OF STATE:

The amounts of reactants and products change

**Chemical Reaction:** some bonds break, and new bonds form

### STATE FUNCTION:

A property that depends only on the state of the system

### PATH FUNCTION:

A property that depends on how a change takes place

### Molar Heat Capacity:

The amount of heat needed to raise the temperature of 1 mol of substance by 1 Kelvin (1K)

### WORK (W):

Energy used to move an object against an opposing force,  $w = Fd$

### FIRST LAW OF THERMODYNAMICS:

$$\Delta E_{\text{sys}} = q_{\text{sys}} + w_{\text{sys}}$$

### Exothermic Process:

If the chemicals release heat, this heat gain raises the temperature of the surroundings

### Endothermic Process:

If the chemicals absorb heat, this heat loss lowers the temperature of the surroundings

$$q_{\text{surroundings}} = -C_{\text{cal}} \Delta T, \quad q = nC_s \Delta T$$

### ENTHALPY:

A thermodynamic quantity whose change equals the heat flow at constant pressure,

$$\Delta H = \Delta E + \Delta(PV)$$

### STANDARD ENTHALPY OF FORMATION ( $\Delta H_f^\circ$ ):

Enthalpy change accompanying the formation of one mole of a chemical substance from pure elements in their most stable forms under standard conditions

### Hess' Law:

The enthalpy change for any overall process is equal to the sum of enthalpy changes for any set of steps that leads from the reactants to the products

### MOLAR HEAT OF SOLUTION:

Measures net energy flow that occurs as substance dissolves

### MOLAR HEAT OF VAPORIZATION:

The heat needed to vaporize one mole of a substance at its normal boiling point

### MOLAR HEAT OF FUSION:

Heat needed to melt one mole of a substance at its normal melting point

### SUBLIMATION:

A phase change in which a solid converts directly to a vapour without passing through the liquid phase,

$$\Delta E_{\text{sub}} = \Delta H_{\text{vap}} - RT_{\text{sub}}$$

### SECOND LAW OF THERMODYNAMICS:

Any spontaneous process increases the disorder of the universe

### ENTROPY:

The state function that provides a quantitative measure of disorder and is symbolized S,

$$\Delta S = \frac{q_r}{T}$$

### Entropy Change of the Universe:

Total entropy change,

$$\Delta S_{\text{universe}} = \Delta S_{\text{system}} + \Delta S_{\text{surroundings}}$$

### REACTION ENTROPIES:

$$\Delta S_{\text{reaction}} = \sum \text{coeff}_p \Delta S^\circ - \sum \text{coeff}_r \Delta S^\circ$$



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# Mcat Chemistry Review Sheet

**Michael Rajnik**



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web elements in the second row of the periodic table place their electrons in the 2n shell as well as the 1n shell for instance lithium li has three electrons two fill the 1 s orbital and the third is placed in the 2 s orbital giving an electron configuration of 1 s 2 2 s 1

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