

Chapter 4 Review Worksheet

Name: _____ Block: _____ Date: _____

1. Which of the following is the smallest particle of an element that can exist by itself?

- A. Ion
- B. Atom**
- C. Molecule
- D. Compound

2. Which of the following correctly matches the subatomic particle with its charge and location in an atom?

	Subatomic Particle	Location	Charge
A	Proton	Nucleus	Neutral
B	Neutron	Nucleus	Positive
C	Electron	Shell	Positive
D	Electron	Shell	Negative

3. Which of the following are responsible for bonding?

- A. Nuclei
- B. Protons
- C. Neutrons
- D. Electrons**

4. How do you calculate the number of neutrons in an atom's nucleus?

- A. Atomic number
- B. Mass number - atomic number**
- C. Mass number + atomic number
- D. Number of electrons + Number of protons

5. Which of the following describes a cation?

I.	Examples include Ca^{2+} and Al^{3+}
II.	A metal atom that has lost electrons
III.	Has an equal number of protons and electrons

- A. I and II only**
- B. I and III only
- C. II and III only
- D. I, II and III

6. Which of the hypothetical elements shown below represents a metal?

- A. Ma
- B. Di
- C. So**
- D. Nh



7. Which of the following does the Bohr model represent?

- A. A neon atom
- B. A sodium atom
- C. A sodium ion**
- D. A fluorine atom



8. Draw Bohr diagrams for the following elements:

a) Nitrogen atom



b) Potassium ion

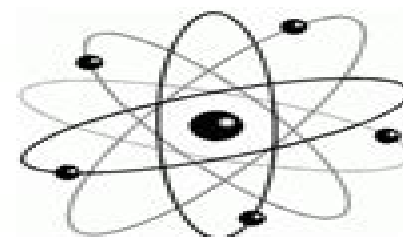


9. Draw Lewis diagrams to show the following chemical bonds:

a) CaCl_2 :



b) CO_2 :



Electrons In Atoms Work Answers

J. Kessler



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Bortz,2003-12-15 Presents the story of the discovery of the first subatomic particle and how it revolutionized understanding of the atom in chemistry and led to the new field of electronics

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