

- Which electron configuration represents an atom in an excited state?
 A) $1s^2 2s^2 2p^6 3p^1$
 B) $1s^2 2s^2 2p^6 3s^2 3p^1$
 C) $1s^2 2s^2 2p^6 3s^2 3p^2$
 D) $1s^2 2s^2 2p^6 3s^2$
- All of the elements in Period 3 have a total of 2 electrons in the
 A) 2s sublevel
 B) 3s sublevel
 C) 2p sublevel
 D) 3p sublevel
- Which atom in the ground state has three unpaired electrons in its outermost principal energy level?
 A) Li
 B) B
 C) N
 D) Ne
- What is the total number of valence electrons in an atom with the electron configuration $1s^2 2s^2 2p^6 3s^2 3p^3$?
 A) 6
 B) 2
 C) 3
 D) 5
- Which orbital notation represents an atom of beryllium in the ground state?
 A)

1s	2s	2p		
<div style="text-align: center;">↑</div>	<div style="text-align: center;">↑</div>	<div style="text-align: center;">↑</div>	<div style="text-align: center;">↑</div>	<div style="text-align: center;"> </div>

 B)

1s	2s	2p		
<div style="text-align: center;">↑↓</div>	<div style="text-align: center;">↑</div>	<div style="text-align: center;">↑</div>	<div style="text-align: center;"> </div>	<div style="text-align: center;"> </div>

 C)

1s	2s	2p		
<div style="text-align: center;">↑↓</div>	<div style="text-align: center;">↑↓</div>	<div style="text-align: center;"> </div>	<div style="text-align: center;"> </div>	<div style="text-align: center;"> </div>

 D)

1s	2s	2p		
<div style="text-align: center;">↑</div>	<div style="text-align: center;"> </div>	<div style="text-align: center;">↑</div>	<div style="text-align: center;">↑</div>	<div style="text-align: center;">↑</div>
- Which element has a completely filled third principal energy level?
 A) Ar
 B) N
 C) Fe
 D) Zn
- Which atom in the ground state has only 3 electrons in the 3p sublevel?
 A) phosphorus
 B) potassium
 C) argon
 D) aluminum
- What is the total number of occupied principal energy levels in a neutral atom of neon in the ground state?
 A) 1
 B) 2
 C) 3
 D) 4

- Which sublevel configuration correctly represents a completely filled third principal energy level?
 A) $3s^2 3p^5 3d^0$
 B) $3s^2 3p^2 3d^{10}$
 C) $3s^2 3p^6 3d^{10}$
 D) $3s^2 3p^6 3d^6$
- Which atom in the ground state contains a partially filled 3p orbital?
 A) argon
 B) calcium
 C) potassium
 D) aluminum
- In the ground state, the atoms of elements in Period 2 all have the same number of
 A) protons
 B) neutrons
 C) 1s electrons
 D) oxidation states
- Which is the electron configuration for a neutral atom in the ground state?
 A) $1s^2 2s^2 3s^1$
 B) $1s^2 2s^2 2p^4 3s^1$
 C) $1s^2 2s^2 2p^6 3p^1$
 D) $1s^2 2s^2 2p^6 3s^1$
- Which represents the electron configuration of an isotope of oxygen in the ground state?
 A) $1s^2 2s^2 2p^1$
 B) $1s^2 2s^2 2p^2$
 C) $1s^2 2s^2 2p^2$
 D) $1s^2 2s^2 2p^4$
- The total number of sublevels in the fourth principal energy level of an atom is
 A) 1
 B) 2
 C) 3
 D) 4
- Which is a possible electron configuration for argon in the excited state?
 A) $1s^2 2s^2 2p^7 3s^2 3p^5$
 B) $1s^2 2s^2 2p^5 3s^2 3p^6$
 C) $1s^2 2s^2 2p^6 3s^1 3p^7$
 D) $1s^2 2s^2 2p^6 3s^2 3p^5 4s^1$
- Which is a correct description of the shape and spatial orientation of the p orbitals in an atom?
 A) All have the same shape but a different orientation.
 B) All have the same shape and the same orientation.
 C) All have a different shape and a different orientation.
 D) All have a different shape but the same orientation.
- A maximum of 6 electrons can occupy
 A) an s orbital
 B) an s sublevel
 C) a p orbital
 D) a p sublevel
- How does the ground state electron configuration of the hydrogen atom differ from that of a ground state helium atom?
 A) Hydrogen has one electron in a higher energy level.
 B) Hydrogen has two electrons in a lower energy level.
 C) Hydrogen contains a half-filled orbital.
 D) Hydrogen contains a completely filled orbital.

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Julia Schneider



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1. Understanding the eBook Electron Distribution Review Worksheet Answers
 - The Rise of Digital Reading Electron Distribution Review Worksheet Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Electron Distribution Review Worksheet Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Electron Distribution Review Worksheet Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Electron Distribution Review Worksheet Answers
 - Personalized Recommendations
 - Electron Distribution Review Worksheet Answers User Reviews and Ratings
 - Electron Distribution Review Worksheet Answers and Bestseller Lists
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 - Highlighting and Note-Taking Electron Distribution Review Worksheet Answers
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 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Electron Distribution Review Worksheet Answers
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 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Electron Distribution Review Worksheet Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Electron Distribution Review Worksheet Answers
 - Setting Reading Goals Electron Distribution Review Worksheet Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Electron Distribution Review Worksheet Answers
 - Fact-Checking eBook Content of Electron Distribution Review Worksheet Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
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