

CHM 101

Electron Structure Worksheet

Objectives:

- Explain the nature of light and how it reveals the electron structure of the atom.
- Write the electron structure of elements
- Relate electron structure to the periodic table

1. Light is a form of what type of energy?
 - Give 2 examples of other forms of this type of energy.
2. What is the meaning of wavelength and frequency of light?
3. White light produces a continuous spectrum. Elements that are excited produce a line spectrum. What is a line spectrum?
4. Different metals emit different colored light in flame tests. This light originates from inside the metal atoms. Explain what was happening inside these atoms that causes the various colors of light. Be specific.
5. What is an orbital?
6. List the orbitals that exist in the following energy levels. Indicate the number of these orbitals.
 $n = 1$, $n = 2$, $n = 3$, $n = 4$
7. Indicate the electron configuration of these elements using the shorthand notation and the orbital diagrams.



For example - helium: $1s^2$, $1s$

Elements: Li, C, Mg, Fe, Kr

8. In this periodic chart, indicate the outer orbitals being filled with electrons in the rectangles.

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9. Why do elements in the same group have similar properties?

Electron Configuration Worksheet Packet

N.B. Singh

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Table of Contents Electron Configuration Worksheet Packet

1. Understanding the eBook Electron Configuration Worksheet Packet
 - The Rise of Digital Reading Electron Configuration Worksheet Packet
 - Advantages of eBooks Over Traditional Books
2. Identifying Electron Configuration Worksheet Packet
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Electron Configuration Worksheet Packet
 - User-Friendly Interface
4. Exploring eBook Recommendations from Electron Configuration Worksheet Packet
 - Personalized Recommendations
 - Electron Configuration Worksheet Packet User Reviews and Ratings

- Electron Configuration Worksheet Packet and Bestseller Lists
- 5. Accessing Electron Configuration Worksheet Packet Free and Paid eBooks
 - Electron Configuration Worksheet Packet Public Domain eBooks
 - Electron Configuration Worksheet Packet eBook Subscription Services
 - Electron Configuration Worksheet Packet Budget-Friendly Options
- 6. Navigating Electron Configuration Worksheet Packet eBook Formats
 - ePub, PDF, MOBI, and More
 - Electron Configuration Worksheet Packet Compatibility with Devices
 - Electron Configuration Worksheet Packet Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Electron Configuration Worksheet Packet
 - Highlighting and Note-Taking Electron Configuration Worksheet Packet
 - Interactive Elements Electron Configuration Worksheet Packet
- 8. Staying Engaged with Electron Configuration Worksheet Packet
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Electron Configuration Worksheet Packet
- 9. Balancing eBooks and Physical Books Electron Configuration Worksheet Packet
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Electron Configuration Worksheet Packet
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Electron Configuration Worksheet Packet
 - Setting Reading Goals Electron Configuration Worksheet Packet
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Electron Configuration Worksheet Packet
 - Fact-Checking eBook Content of Electron Configuration Worksheet Packet
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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