

Chemistry 1020
Electron configurations worksheet

Write the full electron configurations for the following neutral atoms:

1. Ne $1s^2 2s^2 2p^6$
2. Co $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^7$
3. P $1s^2 2s^2 2p^6 3s^2 3p^3$
4. C $1s^2 2s^2 2p^2$
5. Se $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^4$
6. Ag $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^9$

Write the full electron configuration for the following ions and indicate whether the ions are stable or unstable.

7. Be^+ $1s^2 2s^1$ Not a full valence shell = unstable ion
8. Be^{2+} $1s^2 = [\text{He}] = \text{stable ion}$
9. O^{2+} $1s^2 2s^2 2p^2 = \text{unstable ion}$
10. O^{2-} $1s^2 2s^2 2p^6 = [\text{Ne}] = \text{stable ion}$
11. Al^{3+} $1s^2 2s^2 2p^6 = [\text{Ne}] = \text{stable ion}$
12. Al^{3-} $1s^2 2s^2 2p^6 3s^2 3p^4 = \text{unstable ion}$
13. K^+ $1s^2 2s^2 2p^6 3s^2 3p^6 = [\text{Ar}] = \text{stable ion}$
14. K^- $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 = \text{unstable ion}$

Electron Configuration Worksheet And Lots More Answer

S Marginson



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Table of Contents Electron Configuration Worksheet And Lots More Answer

1. Understanding the eBook Electron Configuration Worksheet And Lots More Answer
 - The Rise of Digital Reading Electron Configuration Worksheet And Lots More Answer
 - Advantages of eBooks Over Traditional Books
2. Identifying Electron Configuration Worksheet And Lots More Answer
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Electron Configuration Worksheet And Lots More Answer
 - User-Friendly Interface
4. Exploring eBook Recommendations from Electron Configuration Worksheet And Lots More Answer
 - Personalized Recommendations
 - Electron Configuration Worksheet And Lots More Answer User Reviews and Ratings
 - Electron Configuration Worksheet And Lots More Answer and Bestseller Lists
5. Accessing Electron Configuration Worksheet And Lots More Answer Free and Paid eBooks
 - Electron Configuration Worksheet And Lots More Answer Public Domain eBooks
 - Electron Configuration Worksheet And Lots More Answer eBook Subscription Services
 - Electron Configuration Worksheet And Lots More Answer Budget-Friendly Options

6. Navigating Electron Configuration Worksheet And Lots More Answer eBook Formats
 - ePub, PDF, MOBI, and More
 - Electron Configuration Worksheet And Lots More Answer Compatibility with Devices
 - Electron Configuration Worksheet And Lots More Answer Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Electron Configuration Worksheet And Lots More Answer
 - Highlighting and Note-Taking Electron Configuration Worksheet And Lots More Answer
 - Interactive Elements Electron Configuration Worksheet And Lots More Answer
8. Staying Engaged with Electron Configuration Worksheet And Lots More Answer
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Electron Configuration Worksheet And Lots More Answer
9. Balancing eBooks and Physical Books Electron Configuration Worksheet And Lots More Answer
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Electron Configuration Worksheet And Lots More Answer
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Electron Configuration Worksheet And Lots More Answer
 - Setting Reading Goals Electron Configuration Worksheet And Lots More Answer
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Electron Configuration Worksheet And Lots More Answer
 - Fact-Checking eBook Content of Electron Configuration Worksheet And Lots More Answer
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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