

**7&8 • Atomic Structure & Periodicity****PRACTICE TEST****Rhe: A**

$$\lambda = 2.18 \times 10^{-8} \text{ m}$$

$$R = 1.097 \times 10^7 \text{ m}^{-1}$$

$$\text{mass of an electron} = 9.11 \times 10^{-31} \text{ kg}$$

$$h = 6.626 \times 10^{-34} \text{ J} \cdot \text{s}$$

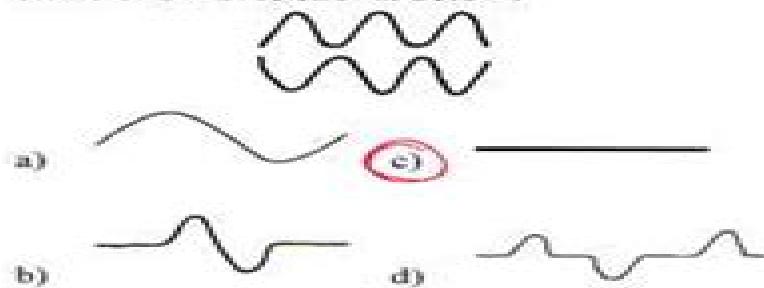
$$c = 3.00 \times 10^8 \text{ m} \cdot \text{s}^{-1}$$

1. What wavelength corresponds to a frequency of  $8.22 \times 10^9 \text{ Hz}$ ?  
 a)  $0.307 \text{ m}$       d)  $0.110 \text{ m}$   
 b)  $0.0365 \text{ m}$        $\lambda = \frac{c}{\nu} = \frac{3.00 \times 10^8 \text{ m/s}}{8.22 \times 10^9 \text{ s}^{-1}} = 0.3649 \text{ m}$   
 c)  $0.122 \text{ m}$
2. A radio station transmits at 110 MHz ( $110 \times 10^6 \text{ Hz}$ ). What wavelength is this radio wave?  
 a)  $3.65 \times 10^{-5} \text{ m}$       c)  $3.81 \times 10^{-5} \text{ m}$   
 b)  $3.30 \text{ m}$       d)  $2.73 \text{ m}$   
 $\lambda = \frac{c}{\nu} = \frac{3.00 \times 10^8 \text{ m/s}}{110 \times 10^6 \text{ s}^{-1}} = 2.73 \text{ m}$
3. Which one of the following is NOT a proper unit for frequency?  
 a) Hz  
 b)  $\text{s}^{-1}$   
 c)  $\text{m} \cdot \text{s}^{-1}$  velocity  
 d)  $\frac{1}{\text{sec}}$
4. Calculate the wavelength of the fourth line in the Balmer series (the visible series) of the hydrogen spectrum.  
 $E_0 = -2.18 \times 10^{-18} \text{ J}$        $E_2 = -2.18 \times 10^{-18} \text{ J}$   
 $\Delta E = 4.84 \times 10^{-19} \text{ J}$   
 $\lambda = \frac{hc}{\Delta E} = \frac{6.626 \times 10^{-34} \text{ J} \cdot \text{s} \cdot 3.00 \times 10^8 \text{ m/s}}{4.84 \times 10^{-19} \text{ J}} = 4.1029 \times 10^{-7} \text{ m}$
5. What is the relationship between the energy of a photon of light and its frequency?  
 a)  $E = \nu$   
 b)  $E = \frac{h}{\nu}$   
 c)  $E = \frac{\nu}{h}$   
 d)  $E = \frac{1}{h\nu}$   
 e)  $E = h\nu$

6. What is the energy needed to raise an electron in the hydrogen atom from the second energy level to the third energy level?  
 a)  $1.52 \times 10^{-19} \text{ J}$   
 b)  $3.63 \times 10^{-19} \text{ J}$   
 c)  $2.18 \times 10^{-19} \text{ J}$   
 $\Delta E = E_3 - E_2 = -2.18 \times 10^{-19} \text{ J} - -2.18 \times 10^{-19} \text{ J} = 2.18 \times 10^{-19} \text{ J}$

- d)  $4.48 \times 10^{-19} \text{ J}$   
 e)  $3.03 \times 10^{-19} \text{ J}$
7. What is the de Broglie wavelength of an electron moving at 80.0% the speed of light?  
 a)  $3.03 \times 10^{-12} \text{ m}$   
 b)  $2.42 \times 10^{-12} \text{ m}$   
 c)  $3.30 \times 10^{11} \text{ m}$   
 d)  $1.59 \times 10^{-25} \text{ m}$

- $\lambda = \frac{h}{mv} = \frac{6.626 \times 10^{-34} \text{ J} \cdot \text{s}}{9.11 \times 10^{-31} \text{ kg} \cdot 8.0 \times 3.00 \times 10^8 \text{ m/s}} = 8.8 \times 10^{-12} \text{ m}$
8. What resultant is expected from the interference of the two waves shown below?



9. Which quantum number determines the subshell occupied by an electron (s, p, d, f, etc.)?  
 a)  $m_s$   
 b)  $\ell$   
 c)  $m_\ell$   
 d)  $m_s$

# Electron Configuration Chem Worksheet 5 6 Answers

**Lingsheng Yao**

**Electron Configuration Chem Worksheet 5 6 Answers:**

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