

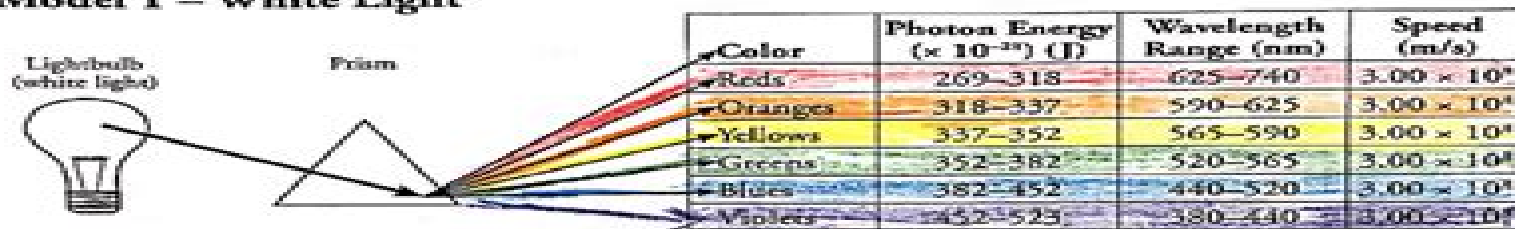
Electron Energy and Light

How does light reveal the behavior of electrons in an atom?

Why?

From fireworks to stars, the color of light is useful in finding out what's in matter. The emission of light by hydrogen and other atoms has played a key role in understanding the electronic structure of atoms. Trace materials, such as evidence from a crime scene, lead in paint or mercury in drinking water, can be identified by heating or burning the materials and examining the color(s) of light given off in the form of bright-line spectra.

Model 1 – White Light



- Trace the arrows in Model 1 and shade in the table with colored pencils where appropriate.
- What happens to white light when it passes through a prism?

The different wavelengths get separated

- Why are the color labels in the table in Model 1 plural (i.e., "Reds" rather than "Red")?

many different shades of a color in a range of wavelength

- Do all colors of light travel at the same speed?

Yes

- Do all colors of light have the same energy? If no, which colors have the highest energy and the least energy, respectively?

*All colors do not have the same energy. Blues + violets highest
Reds + oranges lowest*

- Consider the light illustrated in Model 1.

- Which color corresponds to the longest wavelengths?

Reds

- Which color corresponds to the shortest wavelengths?

Violets

- Write a sentence that describes the relationship between wavelength and energy of light.

As the wavelength gets shorter the energy of light increases

Electron Energy And Light Worksheet Pogil Answers

Harry Dreicer



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Understanding Electrons B. H. Fields,Fred Bortz,2015-07-15 Explore the history of the discovery and the properties of the electron the negatively charged subatomic particle that orbits around an atom s nucleus Energy and Light Peter Lafferty,1993 **On the Calculation of the Electron Energy Spectrum in a Weakly Ionized Gas** N. J. Carron,MISSION RESEARCH CORP SANTA BARBARA CA.,1987 We study an approximate solution of the Boltzmann Equation for the distribution of electrons in a weakly ionized gas in the presence of an electric field E momentum transfer collisions and inelastic energy transfer collisions Except when $E \ll E_0$ we work in the swarm regime where $V \ll v_{drift}$ Low energy electron scattering from light alkali atoms Donald Clifford Gregory,1973 **Experiments on Electron Energy Distributions** Terence Edward KENNY,1968 Photo Electricity Herbert Stanley Allen, Frank Horton, Joseph John Thomson,1913

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