

Electron Configuration Practice

In the space below, write the standard electron configurations of the following elements:

- | | | |
|----|----------|--|
| 1) | nitrogen | $1s^2 2s^2 2p^3$ |
| 2) | copper | $1s^2 2s^2 2p^6 3s^2 3p^6 4s^1 3d^{10}$ |
| 3) | chlorine | $1s^2 2s^2 2p^6 3s^2 3p^5$ |
| 4) | calcium | $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2$ |
| 5) | lead | $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^{14} 5d^{10} 6p^2$ |

In the space below, write the noble gas notation electron configurations of the following elements:

- | | | |
|-----|----------|-----------------------------|
| 6) | barium | $[Xe] 6s^2$ |
| 7) | gold | $[Xe] 4f^{14} 5d^{10} 6s^1$ |
| 8) | platinum | $[Xe] 4f^{14} 5d^9 6s^1$ |
| 9) | xenon | $[Kr] 4d^{10} 5s^2 p^6$ |
| 10) | calcium | $[Ar] 4s^2$ |

Determine what elements are denoted by the following electron configurations:

- | | | |
|-----|--|-------------|
| 11) | $1s^2 2s^2 2p^6 3s^2 3p^3$ | phosphorous |
| 12) | $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^1 4d^7$ | ruthenium |
| 13) | $[Kr] 4d^{10} 5s^2 p^5$ | iodine |
| 14) | $[Xe] 4f^{14} 5d^{10} 6s^2$ | mercury |
| 15) | $[Rn] 7s^2$ | radium |

Determine which of the following electron configurations are not correct:

- | | | |
|-----|---------------------------------|--|
| 16) | $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2$ | $3p^5$ should be a 6 |
| 17) | $1s^3 2s^2 2p^6 3s^2 2p^5$ | 1s only holds 2 and there aren't two "2p" orbitals |
| 18) | $[Cs] 5f^{14} 6d^1 7s^2$ | Cs is not a noble gas |
| 19) | $[Kr] 4d^{10} 5s^2 p^4$ | Correct |
| 20) | $[Ne]$ for Neon | Can't use a noble gas to shortcut itself |

Electron Configuration And Lots More Worksheet Answers

Thomas Griffiths



Electron Configuration And Lots More Worksheet Answers:

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