

Additional Exercises: Galvanic Cell

1. Fill in the table below regarding the redox process in a galvanic cell.

Redox Process	Definition in terms of electron movement	Occurs at which half-cell
Oxidation	<i>Loss of electron</i>	<i>Anode</i>
Reduction	<i>Gain of electron</i>	<i>Cathode</i>

2. For each of the following pair of electrodes, determine the:

- cathode and anode
- half-reactions at the cathode and anode
- cell notation
- the standard cell voltage

- a) Cu & Co

Cathode: Cu, reaction: $\text{Cu}^{2+}(\text{aq}) + 2\text{e}^- \rightarrow \text{Cu}(\text{s})$

Anode: Co, reaction: $\text{Co}(\text{s}) \rightarrow \text{Co}^{2+}(\text{aq}) + 2\text{e}^-$

Cell notation: $\text{Co}(\text{s}) | \text{Co}^{2+}(\text{aq}) || \text{Cu}^{2+}(\text{aq}) | \text{Cu}(\text{s})$

$E^\circ = 0.34\text{V} - (-0.28\text{V}) = 0.62\text{V}$

- b) Zn & Ag

Cathode: Ag, reaction: $\text{Ag}^+(\text{aq}) + \text{e}^- \rightarrow \text{Ag}(\text{s})$

Anode: Zn, reaction: $\text{Zn}(\text{s}) \rightarrow \text{Zn}^{2+}(\text{aq}) + 2\text{e}^-$

Cell notation: $\text{Zn}(\text{s}) | \text{Zn}^{2+}(\text{aq}) || \text{Ag}^+(\text{aq}) | \text{Ag}(\text{s})$

$E^\circ = 0.80\text{V} - (-0.76\text{V}) = 1.56\text{V}$

- c) Mg & Al

Cathode: Al, reaction: $\text{Al}^{3+}(\text{aq}) + 3\text{e}^- \rightarrow \text{Al}(\text{s})$

Anode: Mg, reaction: $\text{Mg}(\text{s}) \rightarrow \text{Mg}^{2+}(\text{aq}) + 2\text{e}^-$

Cell notation: $\text{Mg}(\text{s}) | \text{Mg}^{2+}(\text{aq}) || \text{Al}^{3+}(\text{aq}) | \text{Al}(\text{s})$

$E^\circ = -1.66\text{V} - (-2.37\text{V}) = 0.71\text{V}$

- d) Pb & Zn

Cathode: Pb, reaction: $\text{Pb}^{2+}(\text{aq}) + 2\text{e}^- \rightarrow \text{Pb}(\text{s})$

Anode: Zn, reaction: $\text{Zn}(\text{s}) \rightarrow \text{Zn}^{2+}(\text{aq}) + 2\text{e}^-$

Cell notation: $\text{Zn}(\text{s}) | \text{Zn}^{2+}(\text{aq}) || \text{Pb}^{2+}(\text{aq}) | \text{Pb}(\text{s})$

$E^\circ = -0.13\text{V} - (-0.76\text{V}) = 0.63\text{V}$

- e) Mg & Cu

Cathode: Cu, reaction: $\text{Cu}^{2+}(\text{aq}) + 2\text{e}^- \rightarrow \text{Cu}(\text{s})$

Anode: Mg, reaction: $\text{Mg}(\text{s}) \rightarrow \text{Mg}^{2+}(\text{aq}) + 2\text{e}^-$

Cell notation: $\text{Mg}(\text{s}) | \text{Mg}^{2+}(\text{aq}) || \text{Cu}^{2+}(\text{aq}) | \text{Cu}(\text{s})$

$E^\circ = 0.34\text{V} - (-2.37\text{V}) = 2.71\text{V}$

Electrochemistry Exercise Answers

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