

Chapter 21 • Electrochemistry

(See page 3 for formulas and an answer key)

PRACTICE TEST

1. Which of the following is the correct cell notation for the reaction



- $\text{Cd}^{2+} | \text{Cd} || \text{Hg}_2^{2+} | \text{Hg}$
- $\text{Cd}^{2+} | \text{Hg}_2^{2+} || \text{Cd} | \text{Hg}$
- $\text{Cd} | \text{Cd}^{2+} || \text{Hg}_2^{2+} | \text{Hg}$
- $\text{Cd}^{2+} | \text{Hg} || \text{Hg}_2^{2+} | \text{Cd}$
- $\text{Hg} | \text{Cd} || \text{Hg}_2^{2+} | \text{Cd}^{2+}$

2. Consider an electrochemical cell where the following reaction takes place:



Which of the following is the correct cell notation for this cell?

- $\text{Al} | \text{Al}^{3+} || \text{Sn}^{2+} | \text{Sn}$
- $\text{Al}^{3+} | \text{Al} || \text{Sn} | \text{Sn}^{2+}$
- $\text{Sn} | \text{Sn}^{2+} || \text{Al}^{3+} | \text{Al}$
- $\text{Sn} | \text{Al}^{3+} || \text{Al} | \text{Sn}^{2+}$
- $\text{Al} | \text{Sn}^{2+} || \text{Sn} | \text{Al}^{3+}$

Standard Reduction Potentials at 25°C E° (volts)

$\text{F}_2(\text{g}) + 2\text{e}^- \rightarrow 2\text{F}^-(\text{aq})$	+2.87
$\text{Au}^{3+} + 3\text{e}^- \rightarrow \text{Au(s)}$	+1.50
$\text{Cl}_2(\text{g}) + 2\text{e}^- \rightarrow 2\text{Cl}^-(\text{aq})$	+1.36
$\text{O}_2(\text{g}) + 4\text{H}_3\text{O}^+(\text{aq}) + 4\text{e}^- \rightarrow 6\text{H}_2\text{O(l)}$	+1.23
$\text{Br}_2(\text{l}) + 2\text{e}^- \rightarrow 2\text{Br}^-(\text{aq})$	+1.08
$\text{Ag}^+(\text{aq}) + \text{e}^- \rightarrow \text{Ag(s)}$	+0.80
$\text{Hg}_2^{2+}(\text{aq}) + 2\text{e}^- \rightarrow 2\text{Hg(l)}$	+0.79
$\text{I}_2(\text{s}) + 2\text{e}^- \rightarrow 2\text{I}^-(\text{aq})$	+0.535
$\text{Cu}^{2+}(\text{aq}) + 2\text{e}^- \rightarrow \text{Cu(s)}$	+0.337
$\text{Sn}^{4+}(\text{aq}) + 2\text{e}^- \rightarrow \text{Sn}^{2+}(\text{aq})$	+0.15
$\text{Sn}^{2+}(\text{aq}) + 2\text{e}^- \rightarrow \text{Sn(s)}$	-0.14
$\text{Cd}^{2+}(\text{aq}) + 2\text{e}^- \rightarrow \text{Cd(s)}$	-0.40
$\text{Zn}^{2+}(\text{aq}) + 2\text{e}^- \rightarrow \text{Zn(s)}$	-0.763
$2\text{H}_2\text{O(l)} + 2\text{e}^- \rightarrow \text{H}_2(\text{g}) + 2\text{OH}^-(\text{aq})$	-0.828
$\text{Al}^{3+}(\text{aq}) + 3\text{e}^- \rightarrow \text{Al(s)}$	-1.66
$\text{K}^+(\text{aq}) + \text{e}^- \rightarrow \text{K(s)}$	-2.93
$\text{Li}^+(\text{aq}) + \text{e}^- \rightarrow \text{Li(s)}$	-3.045

3. Given the two half reactions and their potentials, which net reaction is spontaneous?



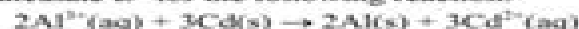
- $\text{Ni(s)} + \text{Mg}^{2+}(\text{aq}) \rightarrow \text{Mg(s)} + \text{Ni}^{2+}(\text{aq})$
- $\text{Ni}^{2+}(\text{aq}) + \text{Mg(s)} \rightarrow \text{Mg}^{2+}(\text{aq}) + \text{Ni(s)}$
- $\text{Ni(s)} + \text{Mg(s)} \rightarrow \text{Mg}^{2+}(\text{aq}) + \text{Ni}^{2+}(\text{aq})$
- $\text{Mg}^{2+}(\text{aq}) + \text{Ni}^{2+}(\text{aq}) \rightarrow \text{Mg(s)} + \text{Ni(s)}$
- $\text{Mg}^{2+}(\text{aq}) + \text{Mg(s)} \rightarrow \text{Ni(s)} + \text{Ni}^{2+}(\text{aq})$

4. Calculate E° for the following reaction:



- +6.00 V
- 3.08 V
- +3.08 V
- +2.78 V
- 2.78 V

5. Calculate E° for the following reaction:



- 2.06 V
- +4.52 V
- +2.06 V
- 4.52 V
- 1.26 V

6. Using data from the reduction potential table and the reaction



calculate the standard reduction potential of the half-reaction



- 1.18 V
- 0.40 V
- 0.40 V
- 1.18 V
- 2.00 V

7. Using data from the reduction potential table, predict which of the following is the best oxidizing agent.

- F_2
- Ag
- Sn^{4+}
- Ag^+
- Al^{3+}

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Minjie Lin



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