#### IGCSE - HONORS - ELECTROCHEMISTRY NOTES AND STUDY GUIDE:

#### 10. Electrolysis of copper - refining copper

Electrodes: metallic copper

Electrolyte: \_\_Copper sulfate Solution

Ions Present: Cu2+ SO4 2- OH- H+



In the Cathode Cu 2+ + 2e- -> Cu

In the Anode Cu -> Cu 2+ + 2e-

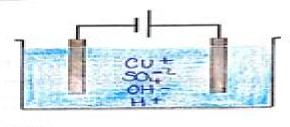
Overall Equation CU 2+ + Cu -> Cu + Cu 2+

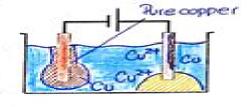
#### Observations:

Cathode: Dork solid deposit (pure copper).....

Anode: Desintegration of anode and yellow deposit below

Uses of copper\_Electrical\_wiring\_capper\_plating\_\_\_\_\_





(show in the graph how it will look like) Why is it necessary to purify Copper? In this way, it will have more conductivity.

What are the elements present in the stag? Aq., Au., Pt.

#### 11. Electroplating with Copper

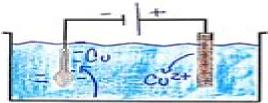
#### Electrodes:

Cathode: Spoon (metal more reactive than Cu)

Anode: Pure copper

Electrolyte: Copper Sulfate or Nitrate

lons Present: CU 21, SO 4 2-, QH -, H+



#### Half equations:

In the Cathode (002++2e -> Cu

In the Anode Cu -> Cu 2+ + 2e -

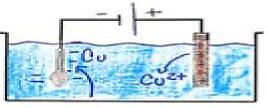
Overall Equation  $C_0^{-2+} + 2e^- + C_0 \longrightarrow C_0 + C_0^{-2+}$ 

#### Observations:

Cathoda: the color of the spoon becomes douber

Anodo: Disintegrate Copper.

Artalia Sanchez



# **Electrochemistry Study Guide Answers**

**Bernice G. Segal, Peter S. Shenkin** 

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