

Electrochemical Cells and Thermodynamics

17

	Shorthand cell designation	Temperature (°C)	E cell (measured)	ΔG (calculated)	K_{eq} (calculated)
1.	Pb Pb Cu	20.2	.492 V	-94 kJ	1601.25
2.	Sn Cu	20.2 20.3	.249 V	-78 kJ	2,198
3.	Sn Pb	20.9	.125 V	-16 kJ	3.645

Show calculations for ΔG° and K_{eq} for an exemplary pair. For Pb | Pb²⁺ || Sn²⁺ | Sn:

Half-cell equation	E half-cell (calculated)
1. $(Pb(s) \rightarrow Pb^{2+} + 2e^-) (Cu^{2+} + 2e^- \rightarrow Cu)$	= .406 V
2. $(Sn \rightarrow Sn^{2+} + 2e^-) (Cu^{2+} + 2e^- \rightarrow Cu)$	= .425 V
3. $(Sn \rightarrow Sn^{2+} + 2e^-) (Pb^{2+} + 2e^- \rightarrow Pb)$	= .0196 V

Effect of Temperature on Cell Potential

Cell designation: E (measured)	Temperature (°C)	Temperature (K)	ΔG (calculated)
Pb Cu: .08 V	25	298	-16,590
Sn Cu: .085 V	35	308	-16,405
Su Pb: .086 V	50	323	-16,212
Pb Cu: .082 V	60	333	-18,120
Su Cu: .081 V	65	338	-18,653
Su Pb: .079 V	70	343	-18,297

ΔS determined from the slope of a plot of ΔG versus T 3.62

ΔF° calculated at 298 K -7.652 (show calculations)

Electrochemical Cells And Thermodynamics Lab Answers

Clifford Lane



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