1. Consider the following reaction.

An equilibrium mixture of this teachon at a certain timp was found to have equilibrium constant (Fe) at this teachon?

K= (0.13) = (0.14) = 136.5

z. Consider the fallowing reaction:

An equilibrium mixture of this teachon at a certain temperature was fund to have [NH₂] = 0.774 M and [NH₂S] + 0.58504. What is the value of the equilibrium constant (FC) at this temperature?

K. [NH₃]CH₂S] - (0.274)(0.385) = 0.0973

3. consider the following reaction:

The reaction between to and Hz is carried but at a specific temperature with within the concentrations of CD = 1-3214 and Hz = 0.4544. At equipment, the concentration of CH30H is 0.1414. Find the equilibrium constant at this temperature.

CO 2HZ CH30H K. (0.14) C -X -2x +X E 0.18 0.17 0.14

4. consider the following reaction:

The reaction of the is carried out at some temperature with an initial concentration of the is only at equilibrium, the concentration of the is only the equilibrium and the equilibrium constant at this temperature.

2 CH4 (2H2 3HZ K- (0.04)10.012)5 - 1.0e-6 1 0.09 Ø Ø (0.082)2 - 1.0e-6 C -2x +x +3x 6 0.082 0.004 0.012M

5. Consider me fallowing reaction

A reaction mixing e contains Duez 1. 0. orner and [NzO4] . 0. 0337 M. CALCHAR QL at the given concentrations

G. The following reaction was carried out in a z.751 reaction vessel at

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Emilie Sanchez

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