

KINETICS Practice Problems and Solutions

Name: _____
 Period: _____ Date: _____

AP Chemistry
 Dr. Mandes

The following questions represent potential types of quiz questions. Please answer each question completely and thoroughly. The solutions will be posted on-line on Monday.

5. Please do #18 in chapter 12 of your text.
- This increases the $[H_2]$, which will increase the rate, but has no effect on k .
 - Due to Arrhenius equation, changing temperature changes the value of k .
 - Catalyst function by lowering the activation energy, so due to Arrhenius equation, changing the activation energy changes the value of k .

6. Consider the reaction: $P_4 + 6 H_2 \rightarrow 4 PH_3$. A rate study of this reaction was conducted at 298 K. The data that were obtained are shown in the table.

$[P_4]$, mol/L	$[H_2]$, mol/L	Initial Rate, mol/(L · s)
0.0110	0.0075	3.20×10^{-4}
0.0110	0.0150	6.40×10^{-4}
0.0220	0.0150	6.39×10^{-4}

- What is the order with respect to:
 P_4 0
 H_2 1
 - Write the rate law for this reaction. $rate = k[H_2]$
 - Determine the value and units of the rate constant, k . plug and chug using the rate law & data from exp't 1 and solving for k , we get $k = 0.0427 s^{-1}$
7. Consider the reaction: $SO_2 + O_3 \rightarrow SO_3 + O_2$. A rate study of this reaction was conducted at 298 K. The data that were obtained are shown in the table.

$[SO_2]$, mol/L	$[O_3]$, mol/L	Initial Rate, mol/(L · s)
0.25	0.40	0.118
0.25	0.20	0.118
0.75	0.20	1.062

- What is the order with respect to:
 SO_2 2
 O_3 0
 - Write the rate law for this reaction. $rate = k[SO_2]^2[O_3]^0$
 - Determine the value and units of the rate constant, k . plug and chug using the rate law & data from exp't 1 and solving for k , we get $k = 2.36 mol/L^{-2} s^{-1}$
8. Consider the following mechanism.
- $$A_2 + B_2 \rightarrow R + C \quad (\text{slow})$$
- $$A_2 + R \rightarrow C \quad (\text{fast})$$
- Write the overall balanced chemical equation. $2 A_2 + B_2 \rightarrow 2 C$
 - Identify any intermediates within the mechanism. R
 - What is the order with respect to each reactant? $A_2: 1^{st}$; $B_2: 1^{st}$

Mastering Chemistry Answer Key Kinetics

WJ Hussar



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