

CK-12 Chemistry - Basic Answer Key

Chapter 7: Chemical Nomenclature

7.1 Ionic Compounds

Lesson Review Questions

Questions

- What is the purpose of an empirical formula?
- Define binary and ternary ionic compounds.
- What is the difference between a monatomic and a polyatomic ion?
- For each of the following ionic compounds, what ions are present and in what ratio?
 - MgBr_2
 - Li_2CO_3
 - $\text{Fe}_2(\text{SO}_4)_3$
- Predict the compound that forms when gallium combines with oxygen. What would the name of this compound be?
- Give the formula for each of the following ionic compounds:

(a) ammonium nitrate	(d) vanadium(III) oxide
(b) cobalt(II) sulfate	(e) barium oxide
(c) nickel(II) cyanide	(f) calcium hypochlorite
- Name the following ionic compounds:
 - MgBr_2
 - (d) KMnO_4
 - Li_2CO_3
 - $(\text{NH}_4)_2\text{S}$
 - KHSO_3
 - CuCl
 - CuCl_2
- Write the correct formulas for the following ionic compounds:
 - Barium chloride
 - Chromium(III) oxide
 - Potassium sulfate
 - Zinc phosphate

Answer:

- The empirical formula is the most straightforward way to describe the structure of a molecule, using a chemical formula that gives the lowest whole number ratio between the ions involved.

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