

## Acids and Bases

**Question 1:** Choose the correct answer for each of the following questions:

- 1.1 The element present in all acids...
- |            |                  |
|------------|------------------|
| A oxygen   | B carbon dioxide |
| C nitrogen | D hydrogen       |
- 1.2 A sample of an acid in a beaker has a pH of 4. An alkali is added to it. What is the pH of the liquid in the beaker?
- |             |                  |
|-------------|------------------|
| A decreases | B stays the same |
| C increases | D neutral        |
- 1.3 Choose the incorrect combination:
- |                                     |   |
|-------------------------------------|---|
| A $\text{HNO}_3$ – Nitric acid.     | B $\text{H}_2\text{SO}_4$ – sulphuric acid. |
| C $\text{HCl}$ – Hydrochloric acid. | D $\text{H}_2\text{CO}_3$ – carbonic acid.  |
- 1.4 The colour of blue litmus paper in an acid is ...
- |          |         |
|----------|---------|
| A yellow | B green |
| C blue   | D red   |
- 1.5 Which substance is always produced in the reaction between hydrochloric acid and lithium?
- |              |                      |
|--------------|----------------------|
| A water      | B hydrogen gas       |
| C oxygen gas | D carbon dioxide gas |
- 1.6 Which substance is always produced in the reaction between hydrochloric acid and sodium hydroxide?
- |              |                |
|--------------|----------------|
| A water      | B hydrogen gas |
| C oxygen gas | D precipitate  |
- 1.7 Which formula represents an acid?
- |                          |                            |
|--------------------------|----------------------------|
| A $\text{KOH}$           | B $\text{KCl}$             |
| C $\text{CH}_3\text{OH}$ | D $\text{CH}_3\text{COOH}$ |
- 1.8 The method of adding an acid to a base is:
- |                  |                 |
|------------------|-----------------|
| A neutralisation | B precipitation |
| C evaporation    | D decomposition |

[8]

**Question 2:** Write the general reaction for an acid with the following compounds:

- 2.1 Base: \_\_\_\_\_
- 2.2 Carbonate: \_\_\_\_\_
- 2.3 Metal: \_\_\_\_\_

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**Victor M. Corman**



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