

Interpretation: Writing Binary Formulas

Example: Andrew Morgan Pre-test: 2 (Date: 2/25/2022)

Using electronegativity as a guidepost and without too many other assumptions, binary writing can be a challenge. When faced with a binary compound, it's always smart to check that both species are neutral. Charge transfer is the number of electrons lost for cations (most cases) and number of electrons gained for anions. The charges are written as the **valence** ions, followed by their symbol and then anions. They also are the magnitude as electrons for anions. Formulas writing are always an art. Sometimes, by simply not understanding.

Directions:

For each case, give the name of the binary compound. Write the formula for each and then name the compound. Remember, the first element is the cation and the second is the anion. The first element is the cation and the second is the anion.

	Cl ⁻	S ²⁻	F ⁻	N ³⁻	O ²⁻	P ³⁻
Mg²⁺	MgCl ₂	MgS ₂	MgF ₂	MgN ₃	MgO ₂	MgP ₃
Ca²⁺	CaCl ₂	CaS ₂	CaF ₂	CaN ₃	CaO ₂	CaP ₃
Cr³⁺	CrCl ₃	CrS ₃	CrF ₃	CrN ₃	CrO ₃	CrP ₃
Na⁺	NaCl	Na ₂ S	NaF	NaN ₃	Na ₂ O	Na ₃ P
Zn²⁺	ZnCl ₂	ZnS	ZnF ₂	ZnN ₃	ZnO	ZnP ₃
Al³⁺	AlCl ₃	Al ₂ S ₃	AlF ₃	AlN	Al ₂ O ₃	AlP

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JG Myers



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