

Electron Configuration Practice

In the space below, write the standard electron configurations of the following elements:

- 1) nitrogen $1s^2 2s^2 2p^3$
- 2) copper $1s^2 2s^2 2p^6 3s^2 3p^6 4s^1 3d^{10}$
- 3) chlorine $1s^2 2s^2 2p^6 3s^2 3p^5$
- 4) calcium $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2$
- 5) lead $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^{14} 5d^{10} 6p^2$

In the space below, write the noble gas notation electron configurations of the following elements:

- 6) barium $[Xe] 6s^2$
- 7) gold $[Xe] 4f^{14} 5d^{10} 6s^1$
- 8) platinum $[Xe] 4f^{14} 5d^9 6s^1$
- 9) xenon $[Kr] 4d^{10} 5s^2 p^6$
- 10) calcium $[Ar] 4s^2$

Determine what elements are denoted by the following electron configurations:

- 11) $1s^2 2s^2 2p^6 3s^2 3p^3$ phosphorous
- 12) $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^1 4d^7$ ruthenium
- 13) $[Kr] 4d^{10} 5s^2 p^5$ iodine
- 14) $[Xe] 4f^{14} 5d^{10} 6s^2$ mercury
- 15) $[Rn] 7s^2$ radium

Determine which of the following electron configurations are not correct:

- 16) $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2$ $3p^5$ should be a 6
- 17) $1s^2 2s^2 2p^6 3s^2 2p^5$ 1s only holds 2 and there aren't two "2p" orbitals
- 18) $[Cs] 5f^{14} 6d^1 7s^2$ Cs is not a noble gas
- 19) $[Kr] 4d^{10} 5s^2 p^4$ Correct
- 20) $[Ne]$ for Neon Can't use a noble gas to shortcut itself

Gpb Chemistry Electron Distribution Answer Key

LL Leslie

A decorative graphic consisting of a light blue horizontal bar with a rounded right end, and a red circular shape partially visible behind it.

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Table of Contents Gpb Chemistry Electron Distribution Answer Key

1. Understanding the eBook Gpb Chemistry Electron Distribution Answer Key
 - The Rise of Digital Reading Gpb Chemistry Electron Distribution Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Gpb Chemistry Electron Distribution Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Gpb Chemistry Electron Distribution Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Gpb Chemistry Electron Distribution Answer Key
 - Personalized Recommendations
 - Gpb Chemistry Electron Distribution Answer Key User Reviews and Ratings
 - Gpb Chemistry Electron Distribution Answer Key and Bestseller Lists
5. Accessing Gpb Chemistry Electron Distribution Answer Key Free and Paid eBooks
 - Gpb Chemistry Electron Distribution Answer Key Public Domain eBooks
 - Gpb Chemistry Electron Distribution Answer Key eBook Subscription Services
 - Gpb Chemistry Electron Distribution Answer Key Budget-Friendly Options

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 - ePub, PDF, MOBI, and More
 - Gpb Chemistry Electron Distribution Answer Key Compatibility with Devices
 - Gpb Chemistry Electron Distribution Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Gpb Chemistry Electron Distribution Answer Key
 - Highlighting and Note-Taking Gpb Chemistry Electron Distribution Answer Key
 - Interactive Elements Gpb Chemistry Electron Distribution Answer Key
8. Staying Engaged with Gpb Chemistry Electron Distribution Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Gpb Chemistry Electron Distribution Answer Key
9. Balancing eBooks and Physical Books Gpb Chemistry Electron Distribution Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Gpb Chemistry Electron Distribution Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Gpb Chemistry Electron Distribution Answer Key
 - Setting Reading Goals Gpb Chemistry Electron Distribution Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Gpb Chemistry Electron Distribution Answer Key
 - Fact-Checking eBook Content of Gpb Chemistry Electron Distribution Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
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 - Integration of Multimedia Elements

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