

7&8 • Atomic Structure & Periodicity

PRACTICE TEST

$$Rhc = A$$

$$A = 2.18 \times 10^{-18} \text{ J}$$

$$h = 6.626 \times 10^{-34} \text{ J}\cdot\text{s}$$

$$R = 1.097 \times 10^7 \text{ m}^{-1}$$

$$c = 3.00 \times 10^8 \text{ m}\cdot\text{s}^{-1}$$

$$\text{mass of an electron} = 9.11 \times 10^{-31} \text{ kg}$$

1. What wavelength corresponds to a frequency of $8.22 \times 10^9 \text{ Hz}$?

$$\lambda = \frac{c}{\nu}$$

- a) 0.307 m d) 0.110 m

- b) 0.0365 m c) 27.4 m

$$\lambda = \frac{3.00 \times 10^8 \text{ m}\cdot\text{s}^{-1}}{8.22 \times 10^9 \text{ s}^{-1}} = 0.036496 \text{ m}$$

2. A radio station transmits at 110 MHz ($110 \times 10^6 \text{ Hz}$). What wavelength is this radio wave?

- a) $3.65 \times 10^{-5} \text{ m}$ c) $3.81 \times 10^{-5} \text{ m}$

- b) 3.30 m d) 2.73 m

$$\lambda = \frac{c}{\nu} = \frac{3.00 \times 10^8 \text{ m}\cdot\text{s}^{-1}}{110 \times 10^6 \text{ s}^{-1}} = 2.73 \text{ m}$$

3. Which one of the following is NOT a proper unit for frequency?

- a) Hz

- b) s^{-1}

- c) $\text{m}\cdot\text{s}^{-1}$

- d) $\frac{1}{\text{sec}}$

velocity

4. Calculate the wavelength of the fourth line in the Balmer series (the visible series) of the hydrogen spectrum.

$$n=6 \rightarrow n=2$$

$$E_6 = -2.18 \times 10^{-18} \text{ J} - E_2 = -2.18 \times 10^{-18} \text{ J}$$

- a) 0.12334 m

- b) 24.373 m

- c) $2.7353 \times 10^{-7} \text{ m}$

- d) $4.1029 \times 10^{-7} \text{ m}$

$$\Delta E = 4.84 \times 10^{-19} \text{ J}$$

$$E = \frac{hc}{\lambda} \quad \lambda = \frac{hc}{E} = 4.1 \times 10^{-7} \text{ m}$$

5. What is the relationship between the energy of a photon of light and its frequency?

- a) $E = \nu$

- b) $E = \frac{h}{\nu}$

- c) $E = h\nu$

- d) $E = \frac{1}{h\nu}$

6. What is the energy needed to raise an electron in the hydrogen atom from the second energy level to the third energy level?

$$\Delta E = E_3 - E_2$$

- a) $1.52 \times 10^{-4} \text{ J}$

- b) $3.63 \times 10^{-19} \text{ J}$

- c) $2.18 \times 10^{-19} \text{ J}$

- d) $4.48 \times 10^{-19} \text{ J}$

- e) $3.03 \times 10^{-19} \text{ J}$

$$-2.18 \times 10^{-18} \text{ J} - (-2.18 \times 10^{-18} \text{ J}) = 3.03 \times 10^{-19} \text{ J}$$

7. What is the de Broglie wavelength of an electron moving at 80.0% the speed of light.

- a) $3.03 \times 10^{-12} \text{ m}$

- b) $2.42 \times 10^{-12} \text{ m}$

- c) $3.30 \times 10^{-11} \text{ m}$

- d) $1.59 \times 10^{-25} \text{ m}$

$$\lambda = \frac{h}{mv} = \frac{6.626 \times 10^{-34} \text{ J}\cdot\text{s}}{9.11 \times 10^{-31} \text{ kg} \times 0.8 \times 3.00 \times 10^8 \text{ m}\cdot\text{s}^{-1}} = 3.03 \times 10^{-12} \text{ m}$$

8. What resultant is expected from the interference of the two waves shown below?



- a)

- b)

- c)

- d)

9. Which quantum number determines the subshell occupied by an electron (s, p, d, f, etc.)?

- a) n

- b) ℓ

- c) m_ℓ

- d) m_s

Gpb Chemistry Answers Atomic Structure

Gabriella Székely



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Mechanisms of Ion Formation and Fragmentation by Fast-atom Bombardment and Matrix-assisted Laser Desorption Ionization Mass Spectrometry Gabriella Székely,1997 **U.S. Government Research Reports** ,1954

Teachers' Journal ,1967 *The Basics of Atoms and Molecules* Tom Jackson,2013-12-15 A fun filled introduction to matter the elements of the periodic table atoms electrons reactions and bonding and radioactivity this volume provides young adults with chemistry examples that reflect their real world interconnections in science Key terms easy experiments and clear illustrations help to guide students through chemical applications A chapter about Niels Bohr and his model for the atom honors his contribution to the understanding of atomic structure and to nuclear fission Tools and techniques such as a scanning tunneling microscope Rutherford s gold foil experiment and a mass spectrometer highlight this instructive text that is aligned to the Common Core Standards **The Atom** Jack Challoner,2018 Until now popular science has relegated the atom to a supporting role in defining the different chemical elements of the periodic table This bold new title places its subject center stage shining the spotlight directly onto the structure and properties of this tiniest amount of anything it is possible to identify The book covers a huge range of topics including the development of scientific thinking about the atom the basic structure of the atom how the interactions between atoms account for the familiar properties of everyday materials the power and mystery of the atomic nucleus and what the mysterious quantum realm of subatomic particles and their interactions can tell us about the very nature of reality Sparkling text banishes an outdated world of dull chemistry as it brightly introduces the reader to what everything is made of and how it all works on the most fundamental level

National Union Catalog ,1980 Includes entries for maps and atlases **National Union Catalog, 1980** ,1981

Index to PB Reports Listed in U.S. Government Research Reports ,1953 *Index to PB Reports Listed in Bibliography of Technical Reports* Technical Information Service, Washington, D.C.,1953 **Index of Conference Proceedings Received** British Library. Lending Division,1983 Chemistry and Atomic Structure Jd Main Smith,2023-07-18 This book is an introduction to the chemistry of atomic structure It covers the basic principles of atomic theory including the nature of atoms chemical bonds and chemical reactions The author provides clear explanations of complex concepts making this book accessible to students and non experts alike Anyone interested in chemistry or atomic theory will find this book a useful resource This work has been selected by scholars as being culturally important and is part of the knowledge base of civilization as we know it This work is in the public domain in the United States of America and possibly other nations Within the United States you may freely copy and distribute this work as no entity individual or corporate has a copyright on the body of the work Scholars believe and we concur that this work is important enough to be preserved reproduced and made generally available to the public We appreciate your support of the preservation process and thank you for being an important part of keeping this knowledge alive and relevant Atomic and Molecular Structure Joel Chaffee,2011-01-15

Learn about the atom what it is the people responsible for helping us understand it and how it affects us in the world today

The HISTORY of the ATOM the PERIODIC TABLE and RADIOACTIVITY Alex Kimmons, 2021-05-15 THE HISTORY OF THE ATOM THE PERIODIC TABLE and RADIOACTIVITY monochrome Chemistry is an interesting and fundamental branch of science because it gives us the chance to explain the secrets of nature What is water What do we use in our cars as fuel What is aspirin What are perfumes made of These kinds of questions and their answers are all part of the world of chemistry There is no industry that does not depend upon chemical substances the petroleum pharmaceuticals garment aircraft steel and electronics industries for example as well as agriculture all utilize the science of chemistry This book helps everyone to understand nature However one does not need to be a chemist or scientist to understand the simplicity within the complexity around us The aim was to write a modern up to date book where students and teachers can get concise information about the structure of substances Sometimes reactions are given in the detailed form but in general excessive detail has been omitted The book is designed to introduce fundamental knowledge in three areas the history of the atom the periodic table and radioactivity We will study the historical development of atomic structure theories the tendencies of elements in periods and groups and the types of emissions and uses of radioactivity *Atomic Theory and Structure of the Atom* T. A. H.

Peacocke, 2013-10-22 Atomic and Nuclear Chemistry Volume 1 Atomic Theory and Structure of the Atom presents the modern ideas of the atomic theory and atomic structure against the background of their historical development Topics covered include the classification of elements atoms and electrons the wave mechanical model of the atom and the determination of atomic weights This volume is comprised of six chapters and begins by discussing the origin of the atomic theory focusing on the role of John Dalton Avogadro's hypothesis and the introduction to the laws of chemical combination The chapters that follow look at the work of the early scientists that led to the development of the periodic table of elements the use of the Avogadro number to determine the actual masses of atoms and molecules and the structure of the atom The essential results of the simple wave mechanical treatment are summarized in the next chapter This book concludes by considering developments in the determination of atomic weights Some brief notes on the character and personality of the great scientists who are mentioned throughout the text are included This book is intended for students and practitioners in the fields of chemistry and physics **ATOMIC STRUCTURE High School Chemistry Curriculum** Bud

Morgan, 2025-07-09 Welcome to our comprehensive guide to Atomic Structure designed specifically for high school students studying chemistry In this book we will explore the fascinating world of atoms the basic building blocks of matter and delve into the complexities of their structure Chapter 1 provides a solid foundation by introducing the concept of atoms and their historical significance in the development of modern chemistry We will examine the key scientists who contributed to our understanding of atomic structure from Democritus to Thomson Rutherford and Bohr Moving on to Chapter 2 we will explore the structure of the atom in more detail including the subatomic particles that make up an atom protons neutrons and

electrons We will also discuss the concept of atomic number atomic mass and isotopes In Chapter 3 we will delve into the arrangement of electrons within an atom s electron cloud including energy levels sublevels and orbital shapes We will also discuss the principles of electron configuration and how it impacts an atom s chemical properties Chapter 4 will focus on the periodic table and how it organizes elements based on their atomic structure We will explore periodic trends such as atomic radius ionization energy and electron affinity and discuss how these trends are related to an element s position on the table In our final chapter we will discuss the significance of atomic structure in various chemical reactions including bonding stoichiometry and the behavior of gases We will also explore the importance of understanding atomic structure in fields such as biochemistry environmental science and materials science Throughout this book we will provide clear explanations helpful illustrations and interactive activities to reinforce key concepts Whether you are a visual learner a hands on learner or a theoretical learner this book is designed to meet your needs We hope that this book will not only enhance your understanding of atomic structure but also inspire you to further explore the wonders of chemistry So let s dive in and unlock the mysteries of the atom together

Atomic Structure and Chemical Reactions Nevin Katz,2007-05 Featuring the comic strip Dr Birdley investigates

Exploration of the Structure of Atom Chandan Sengupta, First Publication April 2019 Revised Publication October 2022 Second Revised Edition July 2023 Third Revised Edition December 2024 Published by Chandan Senguta Printed by IECIT Printing and ublication Services Resource Centre Arabinda Nagar Bankura 722101 W B India Topics Covered Atoms and Molecules Structure of Atom Bonding Mechanism and Chemical Reactions Mechanism of Bonding This book is suitable for students of Class 9 to Class 11 Students aspiring for Pre Medical Entrance Examination can also get adequate support Additional Hard Copies can also be obtained from Chandan Sukumar Sengupta Arabinda Nagar Bankura 722101 WB Write to Us for more materials

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Discovering Atoms Natalie Goldstein,Margaret Christine Campbell,2011-08-15 Provides a history of atoms explores the properties and laws of chemicals and discusses the new subatomic particles

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