

QUESTIONSHEET 1

DYNAMIC EQUILIBRIUM 1

- a) (i) All components of the system are in the **same** phase / **gaseous** phase (1)
- (ii) Both forward and reverse reactions are continuing (1)
but **at the same rate** in opposite directions (1)
- (iii) The overall composition of the mixture (1)
undergoes no net change (1) (as a result of the continuing reactions)
- b) (i) Since this is an equilibrium, all four components are present (1)
 $\text{Cl}_2(\text{aq})$ is also in equilibrium with $\text{Cl}_2(\text{g})$ – hence the smell (1)
- (ii) NaOH neutralises the acids HOCl & HCl forming salts (1)
The removal of products encourages the reaction to go to completion (1)
There is thus no Cl_2 left (1)
- c) (i) *Explanation* Rate at which solid dissolves (1)
is equal to (1)
rate at which ions are deposited (1)
- Equation* $\text{PbCl}_2(\text{s}) + \text{aq} \rightleftharpoons \text{Pb}^{2+}(\text{aq}) + 2\text{Cl}^{-}(\text{aq})$ (1 for species; 1 for state symbols)
- (ii) *Name* Lead(II) chloride (1)
- Explanation* High concentration of Cl^{-} ions from the HCl (1)
displaces equilibrium to the LHS (1)

Frankly Chemistry 20mark Scheme

Judd E. Hollander



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