Writing K_a

 $HC_2H_3O_2(aq) + H_2O(I) \rightleftharpoons H_3O^+(aq) + C_2H_3O_2^-(aq)$

Which of the following is the correct K_a expression for the reaction shown above?

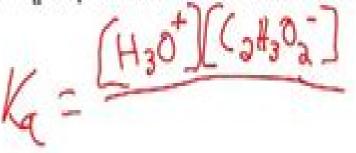
(A)
$$K_a = \frac{[H_3O^+][C_2H_3O_2^-]}{[HC_2H_3O_2][H_2O]}$$

(B)
$$K_a = \frac{[H_3O^+][C_2H_3O_2^-]}{[HC_2H_3O_2]}$$

(C)
$$K_a = \frac{[C_2H_3O_2^-]}{[HC_2H_3O_2]}$$

(D)
$$K_a = \frac{[H_3O^+]}{[H_2O]}$$

(E)
$$K_a = \frac{[HC_2H_3O_2]}{[H_3O^+][C_2H_3O_2^-]}$$



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