

**MARKING SCHEME**  
**SAMPLE PAPER (2024 -25)**  
**CHEMISTRY THEORY**  
**(043)**

| <b>SECTION A</b> |                                                                                                                                                                                                                                                                                                      |            |
|------------------|------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------|------------|
| 1                | (c) reacts with Benzenesulphonyl chloride to form a product that is insoluble in alkali                                                                                                                                                                                                              | 1          |
| 2                | (b) $\text{CH}_3\text{Cl}$<br>The order followed is this $\text{CH}_3\text{I} < \text{CH}_3\text{Br} < \text{CH}_3\text{F} < \text{CH}_3\text{Cl}$ , though F is most electronegative, the bond length is small as compared to C-Cl. Dipole moment is the product of the charge and the bond length. | 1          |
| 3                | (a) (i)-(C), (ii)-(B), (iii)-(A)                                                                                                                                                                                                                                                                     | 1          |
| 4                | (d) 1 =Bromomethane, 2= 2-Bromo-2-methylpropane, 3=2-Bromobutane, 4= 1-Bromobutane<br><br><b>(for visually challenged learners)</b><br>d. 1-Bromobutane                                                                                                                                              | 1<br><br>1 |
| 5                | (c) the order of reaction is zero as the unit of k is $\text{molL}^{-1}\text{s}^{-1}$ . Thus half life = $[\text{R}]_0 / 2k = 4.62 \times 10^{-2} / 2 \times 2.31 \times 10^{-2}$                                                                                                                    | 1          |
| 6                | (b) Benzoic acid and ethanoic acid<br><br>$\text{C}_6\text{H}_5\text{COOCOCH}_3 \xrightarrow{\text{H}_2\text{O}} \text{C}_6\text{H}_5\text{COOH} + \text{CH}_3\text{COOH}$                                                                                                                           | 1          |
| 7                | (b) $\text{X}=[\text{Co}(\text{NH}_3)_4\text{Cl}_2]^+\text{Cl}^-$ , Y= 1:3                                                                                                                                                                                                                           | 1          |
| 8                | (b) Cellulose<br>Starch contains only $\alpha$ glucose, sucrose contains $\alpha$ -D-glucose and $\beta$ -D-fructose glucose, maltose contains $\alpha$ -D-glucose and cellulose is a polymer of $\beta$ -D-glucose.                                                                                 | 1          |
| 9                | (d) $\text{Ti}^{3+} < \text{Cr}^{3+} < \text{Fe}^{2+} < \text{Mn}^{2+}$<br>No. of unpaired electrons : $\text{Ti}^{3+}$ (1), $\text{Cr}^{3+}$ (3), $\text{Fe}^{2+}$ (4) and $\text{Mn}^{2+}$ (5)<br>Paramagnetism depends on the number of unpaired electrons                                        | 1          |
| 10               | (d) It never goes to completion<br>First order reaction $[\text{R}] = [\text{R}_0] e^{-kt}$ If $[\text{R}]=0$ then                                                                                                                                                                                   | 1          |

# K94749 Chemitry Mark Scheme

**Rachel S Tattersall**



**K94749 Chemitry Mark Sheme:**

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