

balance nuclear reaction equations

aluminum-26 undergoes positron emission
potassium-40 undergoes beta decay (electron)
uranium-238 undergoes alpha decay



electron configurations

Write the electron configuration for C.

Write the electron configuration for Cr.

Write the shorthand notation for Bi.



NOMENCLATURE [naming compounds, writing formulas]

Write the formula for each of the following compounds:

- a. Sulfur hexafluoride SF_6
- b. Lithium nitride Li_3N
- c. Chromium (III) carbonate $\text{Cr}_2(\text{CO}_3)_3$
- d. Tin (II) chloride SnCl_2
- e. Ammonium acetate $\text{NH}_4\text{C}_2\text{H}_3\text{O}_2$ (or $\text{CH}_3\text{COONH}_4$)
- f. Mercury (I) chloride Hg_2Cl_2
- g. Potassium bromate KBrO_3
- h. Hydrosulfuric acid $\text{H}_2\text{S(aq)}$
- i. Chloric acid HClO_3
- j. Sulfurous acid H_2SO_3

Name each of the following compounds.

- a. CuSO_4 Copper (II) sulfate
- b. AlF_3 aluminum fluoride
- c. HI hydrogen iodide
- d. NO nitrogen monoxide
- e. $\text{H}_2\text{Se(aq)}$ hydroselenic acid
- f. HNO_3 nitric acid
- g. HNO_2 nitrous acid
- h. NaHSO_4 sodium hydrogen sulfate

Half-life

Tritium ($\text{H}-3$) is a radioactive isotope of hydrogen with a half-life of 12.3 years. How long would it take for a 40.0 g sample to decay down to 1.25 g?

$$40 \xrightarrow{1} 20 \xrightarrow{2} 10 \xrightarrow{3} 5 \xrightarrow{4} 2.5 \xrightarrow{5} 1.25$$

Fe-61 has a half-life of 6.00 min. Of a 100.0 mg sample, how much will remain after 18.0 min?

$$3 \text{ half lives} \quad 100 \xrightarrow{1} 50 \xrightarrow{2} 25 \xrightarrow{3} 12.5 \text{ mg}$$

Average atomic mass

The element copper has naturally occurring isotopes with mass numbers of 63 and 65.

The relative abundance and atomic masses are 69.2% for a mass of 63amu and 30.8% for a mass of 65amu. Calculate the average atomic mass of copper.

$$\begin{array}{rcl} 63 \times .692 & = & 43.596 \rightarrow 44 \\ 65 \times .308 & = & 20.02 \rightarrow 20 \\ + & & \\ \hline & & 63.62 \rightarrow 64 \end{array}$$

K12 Chemistry A Laboratory Guide Answers

Kathleen Armour



K12 Chemistry A Laboratory Guide Answers:

Laboratory Guide to Biochemistry, Enzymology, and Protein Physical Chemistry Marc le Maire, Raymond Chabaud, Guy Hervé, 2012-12-06 The study of a single well chosen substance here aspartate transcarbamoylase can provide an excellent basis for a laboratory course The student is introduced to a variety of scientific ideas and to many experimental and interpretive techniques This enzyme is readily available is relatively stable has an extensive literature and its behavior has many facets substrate inhibition a large change in structure upon homotropic activation by substrates allosteric stimulation by ATP allosteric inhibition by CTP synergistic with VTP positive cooperativity for substrates negative cooperativity for CTP binding and dissociation and reassembly of subunits C and R2 from the holoenzyme CI 5 In addition 3 6 to the known biochemical aspects of these properties the results obtained here can be interpreted in the light of the high resolution X ray diffraction structures of the T and R forms the low angle X ray scattering results and the large number of mutants now available by recombinant DNA methods Future development of this course could also involve part of these methods as well as the carefully chosen experiments described here This approach resembles research more than the approaches one usually finds in biochemical laboratory courses A consistent development of ideas about a single enzyme which shows so many facets in its behavior is sure to hold the interest of the student Moreover one explores a depth and reasons to move forward that are an essential part of research

The Complete Home Learning Sourcebook Rebecca Rupp, 1998 Lists all the resources needed to create a balanced curriculum for homeschooling from preschool to high school level

Illinois Chemistry Teacher ,2007-09 [Nuclear Science Abstracts](#) ,1976-03 *Chemistry of Cement* ,1962

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