

Experiment 9

Electrochemistry I – Galvanic Cell

Introduction: Chemical reactions involving the transfer of electrons from one reactant to another are called *oxidation-reduction* reactions or *redox* reactions. In a redox reaction, two *half-reactions* occur; one reactant gives up electrons (undergoes oxidation) and another reactant gains electrons (undergoes reduction). A piece of zinc going into a solution as zinc ions, with each Zn atom giving up 2 electrons, is an example of an oxidation half-reaction.



The oxidation number of $\text{Zn}(s)$ is 0 and the oxidation number of the Zn^{2+} is +2. Therefore, in this half-reaction, **the oxidation number increases**, which is another way of defining an oxidation. In contrast, the reverse reaction, in which Zn^{2+} ions gain 2 electrons to become Zn atoms, is an example of reduction.



In a reduction there is a **decrease (or reduction) in oxidation number**. Chemical equation representing half-reactions must be both mass and charge balanced. In the half-reactions above, there is one zinc on both sides of the equation. The charge is balanced because the 2+ charge on the zinc ion is balanced by two electrons, $2e^{-}$, giving zero net charge on both sides.

Another example of reduction is the formation of solid copper from copper ions in solution.



In this half-reaction the oxidation number of the aqueous copper is +2, which decreases to 0 for the solid copper, and again charge and mass are balanced. However, no half-reaction can occur by itself. A redox reaction results when an oxidation and a reduction half-reaction are combined to complete a transfer of electrons as in the following example:



The electrons are not shown because they are neither reactants nor products but have simply been transferred from one species to another (from Zn to Cu^{2+} in this case). In this redox reaction, the $\text{Zn}(s)$ is referred to as the *reducing agent* because it causes the Cu^{2+} to be reduced to Cu. The Cu^{2+} is called the *oxidizing agent* because it causes the $\text{Zn}(s)$ to be oxidized to Zn^{2+} .

Any half-reaction can be expressed as a reduction as illustrated in the case where equation (1) can be reversed to equation (2). A measure of the tendency for a reduction to occur is its **reduction potential, E** , measured in units of volts. At standard conditions, 25 °C and concentrations of 1.0 M for the aqueous ions, the measured voltage of the reduction half-reaction is defined as the **standard reduction potential, E°** . Standard reduction potentials

Galvanic Cell Experiment Report

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Report of the President Bryn Mawr College,1896 **Comprehensive Experiments For Materials Science And Engineering** Fei Ye,Chengzhu Liao,Hua Cheng,Jianbo Zhang,Haiou Wang,Yanyan Li,Huili Li,Jing Ming,2023-06-22 The experimental teaching of materials science and engineering MSE is important because the comprehensive applications and the practical knowledge of the professionals are not only an important way for undergraduate students to grasp the knowledge but also to understand the purpose of the study In order to cultivate students ability to solve complex engineering problems more comprehensive experiments should be designed Besides the essential basic experiments in the first few chapters most of the experiments designed in this book are comprehensive hence the title This book breaks the boundaries in the experimental courses of MSE The experiments in this book are modularized into five parts including preliminary exploration of materials science and engineering fundamentals of chemistry and crystallography material properties material preparation and treatment and material applications Besides the experiments the appendices will describe the most relevant aspects of experimental safety error and data presentation in a general way The contents and requirements of the experimental report are suggested At the end of each chapter a list of books journal articles and websites is provided for extended reading on the topics covered in the chapter This book covers the main contents of experimental courses of MSE The experiments cover the forefront of scientific research and the materials industry with appropriate modification It intends to serve as a textbook for undergraduate students and aims to help teachers find a wide enough variety of experiments to construct in an experimental course Solar Energy Update ,1982

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