MASS MOLE STOICHIOMETRY WORKSHEET

Consider the following reaction:

$$N_2 + 2 O_2 \rightarrow N_2 O_4$$

- a. If 15 g of N₂O₄ was produced, how many moles of O₇ were required?
- b. If 4 x 10⁻⁵ moles of oxygen reacted, how many grams of N, were needed?
- 2. Consider the following reaction:

$$K_1PO_4 + Al(NO_1)_1 \rightarrow 3 KNO_1 + AlPO_1$$

- a. What is the mass of potassium nitrate produced when 2.04 moles of potassium phosphate react?
- b. If 5.8 g of aluminum phosphate are formed, how many moles of aluminum nitrate reacted?
- 3. Consider the following reaction:

- a. If you have 5.5 mol of CaC₂, how much C₂H₂ do you get?
- b. How many moles of water are needed when 65 g of CaC; have reacted?
- Hydrogen sulfide gas burns in oxygen to produce sulfur dioxide and water vapor:

What mass of oxygen gas is consumed in a reaction that produces 4.6 mol SO₂?

Mass Mass Stoichiometry Packet Answers

RJ Shavelson

Mass Mass Stoichiometry Packet Answers:

Decoding Mass Mass Stoichiometry Packet Answers: Revealing the Captivating Potential of Verbal Expression

In a time characterized by interconnectedness and an insatiable thirst for knowledge, the captivating potential of verbal expression has emerged as a formidable force. Its ability to evoke sentiments, stimulate introspection, and incite profound transformations is genuinely awe-inspiring. Within the pages of "Mass Mass Stoichiometry Packet Answers," a mesmerizing literary creation penned by way of a celebrated wordsmith, readers embark on an enlightening odyssey, unraveling the intricate significance of language and its enduring effect on our lives. In this appraisal, we shall explore the book is central themes, evaluate its distinctive writing style, and gauge its pervasive influence on the hearts and minds of its readership.

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