

11 Which compound has the greatest percent composition by mass of sulfur?

- way #64
(1) BaS (3) MgS
(2) CaS (4) SrS

way #43
(BA) $2\text{HBr} \rightarrow \text{H}_2 + \text{Br}_2$ (RF)
12 Two molecules of HBr collide and then form H_2 and Br_2 . During the collision, the bonds in the HBr molecules are

- (1) broken as energy is absorbed
(2) broken as energy is released
(3) formed as energy is absorbed
(4) formed as energy is released

way #49
13 Which atom in the ground state has a stable electron configuration?

- (1) carbon (3) neon
(2) magnesium (4) oxygen

way #46
14 Which statement describes a multiple covalent bond?

- (1) Two electrons are shared.
(2) Four electrons are shared.
(3) Two electrons are transferred.
(4) Four electrons are transferred.

way #51
15 The electronegativity difference between the atoms in a molecule of HCl can be used to determine

- (1) the entropy of the atoms
(2) the atomic number of the atoms
(3) the first ionization energy of the atoms
(4) the polarity of the bond between the two atoms

way #5
16 Which two gases can not be broken down by chemical means?

- (1) CO and He (3) Xe and He
(2) CO and NH_3 (4) Xe and NH_3

17 Two substances in a mixture differ in density and particle size. These properties can be used to

- (1) separate the substances
(2) chemically combine the substances
(3) determine the freezing point of the mixture
(4) predict the electrical conductivity of the mixture

18 Which unit is used to express an amount of thermal energy?

- (1) gram (3) joule
(2) mole (4) pascal

19 Under which conditions of temperature and pressure does a real gas behave most like an ideal gas?

- (1) low temperature and low pressure
(2) low temperature and high pressure
(3) high temperature and low pressure
(4) high temperature and high pressure

20 According to the kinetic molecular theory for an ideal gas, all gas particles

- (1) are in random, constant, straight-line motion
(2) are separated by very small distances relative to their sizes
(3) have strong intermolecular forces
(4) have collisions that decrease the total energy of the system

21 Which mathematical expression represents the heat of reaction for a chemical reaction?

- (1) (the heat of fusion) - (the heat of vaporization)
(2) (the heat of vaporization) - (the heat of fusion)
(3) (the potential energy of the products) - (the potential energy of the reactants)
(4) (the potential energy of the reactants) - (the potential energy of the products)

22 At 101.3 kPa and 298 K, a 1.0-mole sample of which compound absorbs the greatest amount of heat as the entire sample dissolves in water?

- (1) LiBr -48.83 (3) NaOH -46.51
(2) NaCl +2.38 (4) $\text{NH}_4\text{Cl} +14.8$

23 For a reaction at equilibrium, which change can increase the rates of the forward and reverse reactions?

- (1) a decrease in the concentration of the reactants
(2) a decrease in the surface area of the products
(3) an increase in the temperature of the system
(4) an increase in the activation energy of the forward reaction

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examines the decay of a Deist universe and the idea of the pre givenness of norms laws and structures as challenged by evolutionary theory and social philosophy It describes the different responses to the collapse of order that given by Christian philosophy scholars who still argue for the idea of a pre given world order and that of other scholars who see this idea of stable creation order and or natural law as redundant and in need of a thorough rethinking It studies the particular role that reformational philosophy has played in the discussion It shows how ever since its inception almost a century ago the concepts of order and law principle structure have been at the heart of this philosophy and that one way to characterise this tradition is as a philosophy of creation order Reformational philosophers have maintained the notion of law as holding for reality This book discusses the questions that have arisen about the nature of such law is it a religious or philosophical concept does law just mean orderliness How does it relate to laws of nature Have they always existed or do they emerge during the process of evolution

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