

Chapter 8: Particle Physics

Outline the results of the gold leaf experiments

1. Alpha particles fired at gold metal foil
2. In a vacuum
3. Most pass straight through/undeflected;
4. Some deflected;
5. Very few reflected/deflected through large angles;
6. Atom mainly empty space;
7. Most did not get near enough to any matter to be deflected;
8. Only a few particles came close enough to charge to be affected;
9. Nucleus has mass much greater than alpha particle mass to cause deflection;
10. Nucleus contains most of the mass;
11. Nucleus very small/tiny;
12. Nucleus charged/positive;
13. Alpha particles are deflected by a charged nucleus that has a very small diameter compared to that of the atom rather than the charge being distributed throughout the atom;
14. most of the mass of the atom is concentrated in the nucleus rather than distributed throughout the atom;

Explain how the observations from the alpha particle scattering experiment led to a new model of the atom

1. Most alpha particles pass through undeflected OR some deflected through a small angle;
2. A very small number are deflected through an angle greater than 90 degrees;
3. This suggests that the alpha particles are deflected by a charged nucleus that has a very small diameter compared to that of the atom rather than the charge being distributed throughout the atom;
4. and that most of the mass of the atom is concentrated in the nucleus rather than distributed throughout the atom;

Role of E and B fields in particle detector

1. Electric field accelerates particles
2. Direction of deflection/force indicates sign of charge;
3. $a = EQ/m$;
4. Magnetic fields provide a centripetal force*;
5. Direction of force/curvature indicates sign of charge;
6. Momentum from radius;
7. $r = p/BQ$;

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RJ Alexander



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