Question	Acceptable Answers		Mark
24(b)(i)	EITHER  C: H: O  mol 64.9: 13.5: 21.6  12: 1: 16  = 5.408: 13.5: 1.35  = 4.006: 10: 1  = 4: 10: 1	(1)	(2)
	use of 74 to show molecular formula is $C_4H_{10}O$ eg $M_r$ is $(4\times12)+(10\times1)+16=74$	(1)	
	OR C atoms = <u>64.9 x 74</u> = 4 100 x 12		
	H atoms = $\frac{13.5 \times 74}{100 \times 1}$ = 10		
	O atoms = 21.6 x 74 = 1 100 x 16		
	This may be done in 2 steps eg C 64.9 x 74 = 48 48 = 4 100 12		
	All 3 correct scores 2 Any 2 correct scores 1		
	OR % C = <u>48 x 100</u> = 64.9		
	% H = <u>10 x 100</u> = 13.5		
	% $O = \frac{16 \times 100}{74} = 21.6$		
	All 3 correct scores (2)		
	Any 2 correct scores (1)		

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Raffaela Di Napoli

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