

Question 33 (4 marks)

A student adds 1.17 g of $\text{Al}(\text{OH})_3(s)$ to 0.500 L of $0.100 \text{ mol L}^{-1} \text{HCl}(aq)$.

4

Calculate the pH of the resulting solution. Assume that the volume of the resulting solution is 0.500 L.

① Write a balanced chemical equation



② Find no. of moles of OH^- and H^+

$$n(\text{Al}(\text{OH})_3) = \frac{m}{M_r} = \frac{1.17}{26.98 + 3(16 + 1.008)} = 0.01499... \text{ mol}$$

$$n(\text{OH}^-) = 3 \times n(\text{Al}(\text{OH})_3) = 0.04499... \text{ mol}$$

$$n(\text{HCl}) = c \times V = 0.5 \times 0.1 = 0.05 \text{ mol}$$

$$n(\text{H}^+) = n(\text{HCl}) = 0.05 \text{ mol}$$

③ Determine which reagent is in excess and how much

As $n(\text{H}^+) > n(\text{OH}^-)$ $\therefore \text{H}^+$ is in excess

$$n(\text{H}^+)_{\text{excess}} = 0.05 - 0.04499... = 0.00500... \text{ mol}$$

④ Calculate the $[\text{H}^+]$ and thus find pH.

$$[\text{H}^+] = \frac{n}{V} = \frac{0.00500...}{0.500} = 0.0100... \text{ mol L}^{-1}$$

$$\text{pH} = -\log_{10} [\text{H}^+] = 2.000 \text{ (3 dp)}$$

Using a calculator for pH

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