

# Unit 6- Mole Conversions (NB p. 49-51)

Vocabulary: Avogadro's number, molar mass, mole ratio, conversion factor, dimensional analysis

Equations & Constants: Avo#:  $6.02 \times 10^{23}$  particles = 1 mole

Practice Problems:

- C 1. The mass of 1 mol of chromium is  
 a. 12 g. ☐ c. 51.996 g.  
 b. 198 g. ☐ d.  $6.02 \times 10^{23}$  g.

$$\frac{1 \text{ mole Cr}}{1 \text{ mole Cr}} \times \frac{g \text{ Cr}}{1 \text{ mole Cr}}$$

- D 2. The mass of a sample containing 3.5 mol of silicon atoms is  
 a. 28 g. ☐ c. 72 g.  
 b. 35 g. ☐ d. 98 g.

$$\frac{3.5 \text{ mol Si}}{1 \text{ mol}} \times \frac{g \text{ Si}}{1 \text{ mol}}$$

- A 3. A prospector finds 39.39 g of gold. She has  
 a.  $1.20 \times 10^{23}$  atoms. ☐ c.  $4.30 \times 10^{23}$  atoms.  
 b.  $2.30 \times 10^{23}$  atoms. ☐ d.  $6.02 \times 10^{23}$  atoms.

- B 4. A sample of tin contains  $3.01 \times 10^{23}$  atoms. The mass of the sample is  
 a. 3.01 g. ☐ c. 72.6 g.  
 b. 59.3 g. ☐ d. 11 g.

$$\frac{3.01 \times 10^{23} \text{ atoms}}{6.02 \times 10^{23} \text{ atoms}} \times \frac{1 \text{ mol}}{1 \text{ mol}} \times \frac{118.71 \text{ g}}{1 \text{ mol}}$$

- B 5. The mass of a sample of nickel is 176.07 g. It contains  
 a.  $1.7607 \times 10^{24}$  atoms. ☐ c.  $5.869 \times 10^{23}$  atoms.  
 b.  $1.806 \times 10^{24}$  atoms. ☐ d.  $5.869 \times 10^{24}$  atoms.

$$\frac{176.07 \text{ g Ni}}{58.693 \text{ g Ni}} \times \frac{1 \text{ mol}}{1 \text{ mol}} \times \frac{6.02 \times 10^{23}}{1 \text{ mol}}$$

- B 6. What is the mass of a 6.267-mol sample of sodium hydroxide?  
 a. 40.00 g ☐ c. 150.4 g  
 b. 250.7 g ☐ d. 6.382 g

$$\text{NaOH} \\ (6.267) \times (40 \text{ g/mol})$$

7. How many atoms of hydrogen are present in 3.18 g of ammonia?  
 a.  $1.12 \times 10^{23}$  ☐ c.  $5.74 \times 10^{24}$   
 b.  $6.33 \times 10^{23}$  ☐ d.  $3.37 \times 10^{23}$

$$\frac{3.18 \text{ g}}{17 \text{ g/mol}}$$

8. How many moles are in 2.3 g of a sample of potassium chlorate?

9. How many carbon atoms are in 9.22 moles of butane,  $\text{C}_4\text{H}_{10}$ ?

Convert grams X to moles X

$$\frac{\text{grams X}}{1 \text{ mole}} \times \frac{1 \text{ mole}}{\text{grams (MM) X}}$$

Convert mole X to grams X

$$\frac{\text{mole X}}{1 \text{ mole X}} \times \frac{\text{grams (MM) X}}{1 \text{ mole X}}$$

Convert atoms of X to moles of X

$$\frac{\text{atoms X}}{6.02 \times 10^{23} \text{ atoms}} \times \frac{1 \text{ mole X}}{1 \text{ mole X}}$$

Convert moles X to atoms of X

$$\frac{\text{moles X}}{1 \text{ mole X}} \times \frac{6.02 \times 10^{23} \text{ atoms}}{1 \text{ mole X}}$$

Convert atoms of X to grams of X

Convert grams of X to atoms of X

How many atoms of X in grams of XY?

How many atoms of X in moles of XY?

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**M Lipman**



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## **Table of Contents Grayslake Chemistry Review Packet Answer Key**

1. Understanding the eBook Grayslake Chemistry Review Packet Answer Key
  - The Rise of Digital Reading Grayslake Chemistry Review Packet Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Grayslake Chemistry Review Packet Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Grayslake Chemistry Review Packet Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Grayslake Chemistry Review Packet Answer Key
  - Personalized Recommendations
  - Grayslake Chemistry Review Packet Answer Key User Reviews and Ratings
  - Grayslake Chemistry Review Packet Answer Key and Bestseller Lists
5. Accessing Grayslake Chemistry Review Packet Answer Key Free and Paid eBooks
  - Grayslake Chemistry Review Packet Answer Key Public Domain eBooks
  - Grayslake Chemistry Review Packet Answer Key eBook Subscription Services
  - Grayslake Chemistry Review Packet Answer Key Budget-Friendly Options
6. Navigating Grayslake Chemistry Review Packet Answer Key eBook Formats
  - ePub, PDF, MOBI, and More
  - Grayslake Chemistry Review Packet Answer Key Compatibility with Devices
  - Grayslake Chemistry Review Packet Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Grayslake Chemistry Review Packet Answer Key
  - Highlighting and Note-Taking Grayslake Chemistry Review Packet Answer Key
  - Interactive Elements Grayslake Chemistry Review Packet Answer Key
8. Staying Engaged with Grayslake Chemistry Review Packet Answer Key

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Grayslake Chemistry Review Packet Answer Key
- 9. Balancing eBooks and Physical Books Grayslake Chemistry Review Packet Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Grayslake Chemistry Review Packet Answer Key
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Grayslake Chemistry Review Packet Answer Key
  - Setting Reading Goals Grayslake Chemistry Review Packet Answer Key
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Grayslake Chemistry Review Packet Answer Key
  - Fact-Checking eBook Content of Grayslake Chemistry Review Packet Answer Key
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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