

Unit Test 1: Matter, Chemical Trends, & Chemical Bonding**Multiple Choice (10 marks) (KU: 3, C: 3, I: 2)**

Identify the choice that best completes the statement or answers the question. Place your answers in the key below.

1. _____ 2. _____ 3. _____ 4. _____ 5. _____
6. _____ 7. _____ 8. _____ 9. _____ 10. _____

- The element found in the periodic table in Group 6 and Period 5 is
a. tungsten, W c. tantalum, Ta
b. molybdenum, Mo d. tellurium, Te
- Which of the following groups is known for having both metallic and non-metallic properties?
a. alkali metals c. transition metals
b. alkali earth metals d. metalloids
- How many electrons does a strontium atom have in its outer shell?
a. 1 c. 7
b. 2 d. 8
- Which particles are responsible for the similar chemical and physical properties of elements in the same group?
a. protons and neutrons c. valence electrons
b. number of energy shells d. isotopes
- Which of the following chemical formulas most likely represents a pure substance containing ionic bonds?
a. Cl_2 c. CO_2
b. SO_2 d. BaO
- Consider the equation $\text{X}_{\text{g}} + e^- \rightarrow \text{X}^- + \text{energy}$. The "energy" term in the equation represents
a. electronegativity c. ionization energy
b. electron affinity d. nuclear charge
- Elements A, B, C, and D (found in Groups 1–17) have atomic radii of $265 \mu\text{m}$, $160 \mu\text{m}$, $185 \mu\text{m}$, and $175 \mu\text{m}$, respectively. Which element will most likely have the lowest ionization energy?
a. A c. C
b. B d. D
- List the following elements from smallest to largest atomic size:
I. barium II. strontium III. bismuth IV. iodine
a. III, I, II, IV c. IV, II, III, I
b. IV, III, II, I d. IV, II, I, III
- Which shape and bond angle is predicted by VSEPR theory for NF_3 ?
a. trigonal pyramid, less than 109.5° c. trigonal planar, 109.5°
b. tetrahedral, 109.5° d. trigonal planar, 120°

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David Baud



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