- 7 Manganese, chromium and ruthenium are all transition elements.
 - (a) Explain what is meant by a transition element.

- (b) MnO_a (aq) and Cr₂O₂ (aq) act as oxidising agents in acidic solution. Both these oxidising agents will oxidise a solution of Sn²⁺ to give a solution of Sn⁴⁺. Solutions containing Sn²⁺ and solutions containing Sn⁴⁺ are colourless.
 - (ii) Describe the colour change seen when an excess of Sn2*(ag) is added separately to
 - dilute acidified MnO_{*} (aq)

from ______ to

dilute acidified Cr₂O₂² (aq).

(ii) Write an equation for the reaction between Sn^{2*}(ag) and acidified Cr₂O₂*(ag).

[1]

(c) Ruthenium, Ru, forms complex ions. In one such complex ion, X, the ruthenium ion has a co-ordination number of six. Each complex ion X contains one Ru³⁺ ion, one Cl⁻ ion, one SO₂ molecule and the remaining ligands are NH₃ molecules.

The SO₂ molecule acts as a monodentate ligand and is attached to the Ru²⁺ ion via the sulfur atom. **X** exists in two isomeric forms.

(i) State what is meant by a co-ordination number of six.

(iii) State the formula of X. Include its charge.

[1]

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