

## Unit 2: Essentials of Geometry

### Lesson 2: All About Angles (Adding and Bisecting)

objectives

key

Today we will learn:

- ✱ Angle Addition Postulate
- ✱ Properties of an Angle Bisector and congruent angles

Piece + Piece = whole

LESSON SUMMARY

### Angle Addition Postulate

- If point B lies in the interior of angle AOC, then

$$m\angle AOB + m\angle BOC = m\angle AOC$$

$$P + P_A = W$$



- \* To find the whole (red) angle → add the pieces
- \* To find any "piece" subtract the given one from the "whole"

### Angle Bisector and Congruence

↓  
"splits" angles into 2 equal parts

↓ means "same measure"



→ OP is bisector  
OR  $\angle AOB$



Congruent Angles

$\angle A \cong \angle B$   
Because

$$m\angle A = m\angle B$$

$\cong \Rightarrow$   
congruent

# Geometry Hs Mathematics Key Unit 2 Lesson 2

**S Nieto**



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web answer  $\text{Na}^+ \text{aq}$  and  $\text{SO}_4^{2-} \text{aq}$  identify the chemical equation the spectator ions  $\text{KOH} \text{aq}$   $\text{HNO}_3$   $\text{KNO}_3 \text{aq}$   $\text{H}_2\text{O}$   $\text{K}^+ \text{aq}$   $\text{OH}^- \text{aq}$   $\text{H}^+ \text{aq}$   $\text{NO}_3^- \text{aq}$   $\text{K}^+ \text{aq}$   $\text{NO}_3^- \text{aq}$   $\text{H}_2\text{O}$  answer  $\text{K}^+ \text{aq}$  and  $\text{NO}_3^- \text{aq}$  identify the chemical equation the spectator ions  $\text{Pb(NO}_3)_2 \text{aq}$   $2\text{KI} \text{aq}$   $\text{PbI}_2 \text{s}$   $2\text{KNO}_3 \text{aq}$

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web these ions are called spectator ions because they do not participate in the actual reaction canceling the spectator ions gives the net ionic equation which shows only those species that participate in the chemical reaction  $2\text{Ag}^+ \text{aq}$   $\text{CrO}_4^{2-} \text{aq}$   $\rightarrow \text{Ag}_2\text{CrO}_4 \text{s}$  label 4 2 3

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web chemical equation in which all dissolved ionic reactants and products including spectator ions are explicitly represented by formulas for their dissociated ions molecular equation chemical equation in which all reactants and products are

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web jan 29 2020 the ionic form of this reaction is  $2\text{Na}^+ \text{aq}$   $2\text{Cl}^- \text{aq}$   $\text{Cu}^{2+} \text{aq}$   $\text{SO}_4^{2-} \text{aq}$   $2\text{Na}^+ \text{aq}$   $\text{SO}_4^{2-} \text{aq}$   $\text{CuCl}_2 \text{s}$  the sodium ions and sulfate ion are the spectator ions in this reaction they appear unchanged in both the product and reactant side of the equation

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web mar 11 2014 solution ionic equation  $\text{H}^+ \text{aq}$   $\text{Cl}^- \text{aq}$   $\text{Na}^+ \text{aq}$   $\text{HCO}_3^- \text{aq}$   $\text{Na}^+ \text{aq}$   $\text{Cl}^- \text{aq}$   $\text{H}_2\text{O}$   $\text{l}$   $\text{CO}_2 \text{g}$  the spectator ions are  $\text{Cl}^-$  and  $\text{Na}^+$  net ionic equation  $\text{H}^+ \text{aq}$   $\text{Cl}^- \text{aq}$   $\text{Na}^+ \text{aq}$   $\text{HCO}_3^- \text{aq}$   $\text{Na}^+ \text{aq}$   $\text{Cl}^- \text{aq}$   $\text{H}_2\text{O}$   $\text{l}$   $\text{CO}_2 \text{g}$   $\text{H}^+ \text{aq}$   $\text{HCO}_3^- \text{aq}$   $\text{H}_2\text{O}$   $\text{l}$   $\text{CO}_2 \text{g}$  here is a video on spectator ions

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web the six  $\text{Na}^+$  ions and the six  $\text{Cl}^-$  ions that appear on both sides of the equation are spectator ions that can be canceled to give the net ionic equation  $\text{Ba}^{2+} + 2\text{PO}_4^{3-} \rightarrow \text{Ba}_3(\text{PO}_4)_2 \text{ s}$

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web these spectator ions whose presence is required to maintain charge neutrality are neither chemically nor physically changed by the process and so they may be eliminated from the equation to yield a more succinct representation called a

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web net ionic equations an equation for a reaction in solution showing only those particles that are directly involved in the chemical change spectator ions an ion that is not directly involved in a chemical reaction an ion that does not change oxidation number or composition during a reaction products

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web the equation below is the final net ionic equation after elimination of the spectator ions which yields water  $\text{H}^+ \text{ aq} + \text{OH}^- \text{ aq} \rightarrow \text{H}_2\text{O l}$  the spectator ions in  $\text{KOH aq}$  and  $\text{HNO}_3 \text{ aq}$  are  $\text{K}^+ \text{ aq}$  and  $\text{NO}_3^- \text{ aq}$  how to identify and list the

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reactions and products  $\text{Pb(NO}_3)_2(\text{aq}) + 2\text{KCl}(\text{aq}) \rightarrow \text{PbCl}_2(\text{s}) + 2\text{KNO}_3(\text{aq})$  complete ionic equation an equation written with all soluble strong electrolytes shown as ions

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web the  $\text{Ag}^+(\text{aq})$  and  $\text{Cl}^-(\text{aq})$  ions become  $\text{AgCl}(\text{s})$  but the  $\text{Na}^+(\text{aq})$  ions and the  $\text{NO}_3^-(\text{aq})$  ions stay as  $\text{Na}^+(\text{aq})$  ions and  $\text{NO}_3^-(\text{aq})$  ions these two ions are examples of spectator ions ions that do nothing in the overall course of a chemical reaction

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web aug 8 2022 a spectator ion is an ion that does not take part in the chemical reaction and is found in solution both before and after the reaction the net ionic equation is the chemical equation that shows only those elements compounds and ions that are directly involved in the chemical reaction

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web jul 26 2017 answer the spectator ions are explanation in the net ionic equations we are not include the spectator ions in the equations spectator ions the ions present on reactant and product side 16 used to grind chemicals to powder tmraor nda stlepe star 4 4 5 heart 181 verified verified answer food

**10826 ap chemistry chemical equations worksheet loudoun** - Jul 30 2022

web the new format requires students to write balanced chemical equations showing only the reacting substances excluding spectator ions for three different sets of reactants and to answer a short question requiring no calculator about each reaction the questions are intended to test the students understanding of the meaning of the reactions

*spectator ions learn their role in reactions examples* - Mar 26 2022

web jun 2 2023 the spectator ions are cut out from a reaction hence the net ionic equation for the above reaction would be  $2\text{Cl}^-(\text{aq}) + \text{Cu}^{2+}(\text{aq}) \rightarrow \text{CuCl}_2(\text{s})$   $2\text{Cl}^-(\text{aq}) + \text{Cu}^{2+}(\text{aq}) \rightarrow \text{CuCl}_2(\text{s})$  these spectator ions are not actively involved in the reaction the existence of spectator ions affects the debye length