



Name: _____

Period: _____

Date: _____

Periodic Table and Periodic Trends Review

1. Why do elements in the same group have similar properties? Same number of valence electrons.
2. Which 3 elements would you assume to have similar properties? O, He, S, Ca, P, Se O, S, Se.
 - Why did you choose those three? They are in the same Group
3. Which elements reside in period 2? Lithium, Beryllium, Boron, Carbon, Nitrogen, Oxygen, Fluorine, Neon.
4. T T or F: Group 1 elements are referred to as alkali metals.
5. F T or F: Calcium is halogen.
6. T T or F: Tungsten is a transition metal.
7. Metals prefer to (gain / lose) electrons and become (cations / anions).
8. Aluminum will (gain / lose) electrons and become a (cation / anion).
9. Chlorine will (gain / lose) electrons and become a (cations / anion).
10. F T or F: Bromine is considered to be a good conductor of electricity.
11. F T or F: Aluminum is a metalloid.
12. T T or F: Silicon is a metalloid.
13. T T or F: Antimony would be considered to have both metal and nonmetal properties.
14. T T or F: Barium has 2 valence electrons.
15. F T or F: Phosphorus has 7 valence electrons
16. Which element has an electron configuration ending in $3p^6$? Argon
17. Which block (s,p,d,f) are the following found?
 - Sodium s
 - Iron d
 - Silicon p
 - Uranium f
18. Which block (s,p,d,f) contains the actinide series? f
19. In general, the electronegativity of an element (increases/decreases) as you move across a period and (increases/decreases) as you move down a group.
20. In general, the atomic radius of an element (increases/decreases) as you move across a period and (increases/decreases) as you move down a group.
21. Identify the period and group of the element with the electron configuration of $[\text{Ne}] 3S^2 3P^4$ Period 3 – Group 16

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A Gutmann



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