

Model 2 – Ground State Orbital Diagrams and Electron Configurations



Hydrogen
 $1s^1$



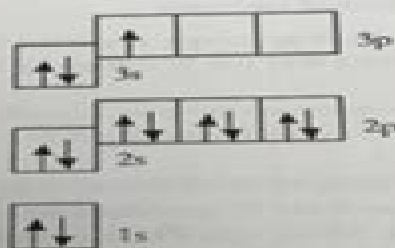
Carbon
 $1s^2 2s^2 2p^2$



Oxygen
 $1s^2 2s^2 2p^4$



Sodium
 $1s^2 2s^2 2p^6 3s^1$



Aluminum
 $1s^2 2s^2 2p^6 3s^2 3p^1$




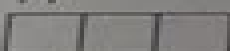
Phosphorus
 $1s^2 2s^2 2p^6 3s^2 3p^3$

6. Examine the orbital diagrams and electron configurations in Model 2. Match each symbol below with its meaning.

___ a. 

___ b. 

___ c. 

___ d. 

___ e. $1s^2 2s^2 2p^4$

I. Single electron

II. Pair of electrons with opposite spins

III. Atomic orbital (region of space where an electron is likely to be found)

IV. Sublevel (set of orbitals having equivalent energy)

V. Electron configuration

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Dale Rich



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