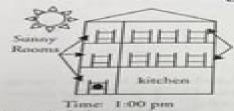
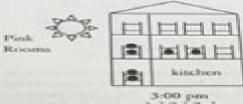
Learning Target - I understand that, when an atom is in ground state, that its electrons occupy the lowestenergy levels, sublevels, and orbitals.

Success Criteria - I can use the periodic table to determine the electron configuration for ground-state atoms.

Model 1 - The Boarding House



Manager's Code:



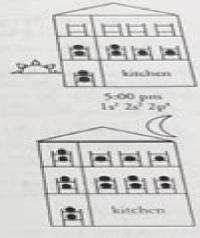
3:00 pm 1s² 2s² 2p³



Time: 7:00 pm Manager's Code: 1s 2s 2p 3s'



9:00 pm 1st 2st 2pt 3st 3pt



11:00 pm 1s' 2s' 2p' 3s' 3p'

1. Examine the boarding house diagrams in Model 1. Match each symbol below with its most likely meaning.

I. Bunk bed for boarders

II. Manager's code for the number of boarders in the house and their room assignments.

c. 1s2 2s2 2p2 3s2

III. Boarder

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Teague, 2014-06-11 In AMEC chapter one you will find a structure for O2 supported by both bond order and unpaired electron data Previous attempts at providing an electronic structure for O2 fails due to bond order number of unpaired electrons or oxygen atoms indicating five orbitals although four is the limit Herein you will find the newly described LONE P ORBITAL Pi BOND which is the solution not only for O2 but also for NO and OF uncharacterized The LPOP bond system is possible due to the small size and high electronegative character of the N O and F atoms which makes it related to H Bonding Diatomic structures involving row two periodic table elements LiBe Be2 BeB B2 BC C2 CN N2 and F2 are also presented All structures are shown being formed by two methods MRAE Most Recently Added Electron and LCAO Linear Combination of Atomic Orbitals Chapter one also includes a Section III MRAE Mechanistic Sequence For NO NO O2 OF F2 and F2 F F0 Section IV Photographs of NO and O2 molecular models Section V MRAE MO Energy Level diagrams for the diatomic structures listed above and Section VI a MRAE MO Energy Level Diagrams Template which is based on the cascade of changes in electronic distribution when an electron is forced onto a previous structure MRAE such as N2 going to N2 and NO going to NO Chapter 2 includes sections on hybridization and mechanism as it relates to polyatomic species The following sequences are given C H CH H CH2 H CH3 H CH4 saturated and tetrahedral arrangement and N H NH H NH2 H NH3 H NH4 also saturated and tetrahedral arrangement The emphasis is to show trend from atomic structure through sp sp2 and sp3 hybridization Chapter 3 is composed of selections from Atomic and Molecular Configuration 1987 Included is MRAE AUFBAU TEMPLATE of the periodic table elements expanded to two pages It is informative to compare the MRAE concept for atoms to the MRAE method for diatomic species in Chapter one In both cases the key is the changes in electronic structure when an electron is added

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