

8 • Electron Configurations & Periodicity

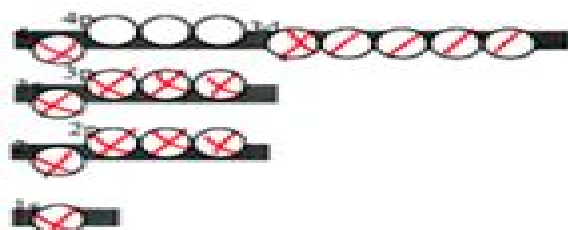
WRITING ELECTRON CONFIGURATIONS

For each given element, fill in the orbital diagram and then write the electron configuration for the element.

1.	2.	3.	4.	5.	6.
Element: Ar # of e ⁻ s: <u>18</u>	Element: Mg # of e ⁻ s: <u>12</u>	Element: N # of e ⁻ s: <u>7</u>	Element: Li # of e ⁻ s: <u>3</u>	Element: P # of e ⁻ s: <u>15</u>	Element: Cl # of e ⁻ s: <u>17</u>

Write the electron configurations of each of these in **long form** and **short form**:

- Ar $1s^2 2s^2 2p^6 3s^2 3p^4$
Ar $[Ne] 3s^2 3p^4$
- Mg $1s^2 2s^2 2p^6 3s^2$
Mg $[Ne] 3s^2$
- N $1s^2 2s^2 2p^3$
N $[He] 2s^2 2p^3$
- Li $1s^2 2s^1$
Li $[He] 2s^1$
- P $1s^2 2s^2 2p^4 3s^2 3p^3$
P $[Ne] 3s^2 3p^3$
- Cl $1s^2 2s^2 2p^6 3s^2 3p^5$
Cl $[Ne] 3s^2 3p^5$



- Fill in the orbital diagram for the element, Fe, and write the electron configuration of Fe in the long and short form.



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Table of Contents Electron Configuration 5 6 Answer Key

1. Understanding the eBook Electron Configuration 5 6 Answer Key
 - The Rise of Digital Reading Electron Configuration 5 6 Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Electron Configuration 5 6 Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Electron Configuration 5 6 Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Electron Configuration 5 6 Answer Key
 - Personalized Recommendations
 - Electron Configuration 5 6 Answer Key User Reviews and Ratings
 - Electron Configuration 5 6 Answer Key and Bestseller Lists

5. Accessing Electron Configuration 5 6 Answer Key Free and Paid eBooks
 - Electron Configuration 5 6 Answer Key Public Domain eBooks
 - Electron Configuration 5 6 Answer Key eBook Subscription Services
 - Electron Configuration 5 6 Answer Key Budget-Friendly Options
6. Navigating Electron Configuration 5 6 Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Electron Configuration 5 6 Answer Key Compatibility with Devices
 - Electron Configuration 5 6 Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Electron Configuration 5 6 Answer Key
 - Highlighting and Note-Taking Electron Configuration 5 6 Answer Key
 - Interactive Elements Electron Configuration 5 6 Answer Key
8. Staying Engaged with Electron Configuration 5 6 Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Electron Configuration 5 6 Answer Key
9. Balancing eBooks and Physical Books Electron Configuration 5 6 Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Electron Configuration 5 6 Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Electron Configuration 5 6 Answer Key
 - Setting Reading Goals Electron Configuration 5 6 Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Electron Configuration 5 6 Answer Key
 - Fact-Checking eBook Content of Electron Configuration 5 6 Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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