

Additional Exercises: Galvanic Cell

1. Fill in the table below regarding the redox process in a galvanic cell.

Redox Process	Definition in terms of electron movement	Occurs at which half-cell
Oxidation	<i>Loss of electron</i>	<i>Anode</i>
Reduction	<i>Gain of electron</i>	<i>Cathode</i>

2. For each of the following pair of electrodes, determine the:

- cathode and anode
- half-reactions at the cathode and anode
- cell notation
- the standard cell voltage

a) Cu & Co

Cathode: Cu, reaction: $\text{Cu}^{2+}(\text{aq}) + 2\text{e}^- \rightarrow \text{Cu}(\text{s})$

Anode: Co, reaction: $\text{Co}(\text{s}) \rightarrow \text{Co}^{2+}(\text{aq}) + 2\text{e}^-$

Cell notation: $\text{Co}(\text{s}) | \text{Co}^{2+}(\text{aq}) || \text{Cu}^{2+}(\text{aq}) | \text{Cu}(\text{s})$

$E^\circ = 0.34\text{V} - (-0.28\text{V}) = 0.62\text{V}$

b) Zn & Ag

Cathode: Ag, reaction: $\text{Ag}^+(\text{aq}) + \text{e}^- \rightarrow \text{Ag}(\text{s})$

Anode: Zn, reaction: $\text{Zn}(\text{s}) \rightarrow \text{Zn}^{2+}(\text{aq}) + 2\text{e}^-$

Cell notation: $\text{Zn}(\text{s}) | \text{Zn}^{2+}(\text{aq}) || \text{Ag}^+(\text{aq}) | \text{Ag}(\text{s})$

$E^\circ = 0.80\text{V} - (-0.76\text{V}) = 1.56\text{V}$

c) Mg & Al

Cathode: Al, reaction: $\text{Al}^{3+}(\text{aq}) + 3\text{e}^- \rightarrow \text{Al}(\text{s})$

Anode: Mg, reaction: $\text{Mg}(\text{s}) \rightarrow \text{Mg}^{2+}(\text{aq}) + 2\text{e}^-$

Cell notation: $\text{Mg}(\text{s}) | \text{Mg}^{2+}(\text{aq}) || \text{Al}^{3+}(\text{aq}) | \text{Al}(\text{s})$

$E^\circ = -1.66\text{V} - (-2.37\text{V}) = 0.71\text{V}$

d) Pb & Zn

Cathode: Pb, reaction: $\text{Pb}^{2+}(\text{aq}) + 2\text{e}^- \rightarrow \text{Pb}(\text{s})$

Anode: Zn, reaction: $\text{Zn}(\text{s}) \rightarrow \text{Zn}^{2+}(\text{aq}) + 2\text{e}^-$

Cell notation: $\text{Zn}(\text{s}) | \text{Zn}^{2+}(\text{aq}) || \text{Pb}^{2+}(\text{aq}) | \text{Pb}(\text{s})$

$E^\circ = -0.13\text{V} - (-0.76\text{V}) = 0.63\text{V}$

e) Mg & Cu

Cathode: Cu, reaction: $\text{Cu}^{2+}(\text{aq}) + 2\text{e}^- \rightarrow \text{Cu}(\text{s})$

Anode: Mg, reaction: $\text{Mg}(\text{s}) \rightarrow \text{Mg}^{2+}(\text{aq}) + 2\text{e}^-$

Cell notation: $\text{Mg}(\text{s}) | \text{Mg}^{2+}(\text{aq}) || \text{Cu}^{2+}(\text{aq}) | \text{Cu}(\text{s})$

$E^\circ = 0.34\text{V} - (-2.37\text{V}) = 2.71\text{V}$

Electrochemistry Exercise Answers

IM Harris



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Chemistry: The Central Science Theodore L. Brown, H. Eugene LeMay Jr., Bruce E. Bursten, Catherine Murphy, Patrick Woodward, Steven Langford, Darius Sagatys, Adrian George, 2013-10-04 If you think you know the Brown LeMay Bursten Chemistry text think again In response to market request we have created the third Australian edition of the US bestseller Chemistry The Central Science An extensive revision has taken this text to new heights Triple checked for scientific accuracy and consistency this edition is a more seamless and cohesive product yet retains the clarity innovative pedagogy functional problem solving and visuals of the previous version All artwork and images are now consistent in quality across the entire text And with a more traditional and logical organisation of the Organic Chemistry content this comprehensive text is the source of all the information and practice problems students are likely to need for conceptual understanding development of problem solving skills reference and test preparation **Modern Aspects of Electrochemistry 40** Ralph E.

White, 2010-04-28 This volume begins with a tribute to Dr Brian E Conway by Dr John O M Bockris which is followed by six chapters The topics covered are state of the art Polymer Electrolyte Membrane PEM fuel cell bipolar plates use of graphs in electrochemical reaction networks nano materials in lithium ion batteries direct methanol fuel cells two chapters and the last chapter presents simulation of polymer electrolyte fuel cell catalyst layers David and Valerie Bloomfield begin the first chapter with a discussion of the difficulties encountered when confronting bipolar plate development and state that the problems stem from the high corrosive nature of phosphoric acid The water problems are mitigated but the oxidation problems increase Bipolar plates are still not cheap reliable or durable In Chapter 2 Thomas Z Fahidy reviews analysis of variance ANOVA and includes one way two way three way classification and Latin squares observation methods He moves on to a discussion of the applications of the analysis of covariance ANCOVA and goes over certain variables such as velocity velocity and pressure drop and product yields in a batch and flow electrolyzer His conclusion is that proper statistical techniques are time savers which can save the experimenter and the process analyst considerable time and effort in trying to optimize the size of statistically meaningful experiments Electrochemical and Metallurgical Industry Eugene Franz

Roeber, Howard Coon Parmelee, 1906 Chemical Principles Steven S. Zumdahl, 1995 The Study Guide reflects the unique problem solving approach taken by the Chemical Principles text The new edition of the Study Guide includes many new worked out examples *Physical Chemistry for JEE Advanced: Part 1, 3E (Free Sample)* K. S. Verma, 2022-05-19 Physical Chemistry for JEE Advanced Part 1 a Cengage Exam Crack Series product is designed to help aspiring engineers focus on the subject of physical chemistry from two standpoints To develop their caliber aptitude and attitude for the engineering field and profession To strengthen their grasp and understanding of the concepts of the subjects of study and their applicability at the grassroots level Each book in this series approaches the subject in a very conceptual and coherent manner While its illustrative solved examples facilitate easy mastering of the concepts and their applications an array of solved problems exposes the students to a variety of questions that they can expect in the examination The coverage and features of this series of books make it highly useful for all those preparing for JEE Main and Advanced and aspiring to become engineers

Solved Problems in Electrochemistry for Universities and Industry Dominique L. Piron, 2004 *Physical Chemistry* David S. Eisenberg, Donald M. Crothers, 1979 Modern Aspects of Electrochemistry John O'M. Bockris, 1954
Indian Army Acc Entrance Exam Arihant Experts, 2018-04-20 Army Cadet College ACC is a wing of Indian Military Academy responsible for providing three years of an intense training to the Indian Soldiers to transform them into Officer for Indian Army Conducting the Indian Army Entrance Exam twice in a year it selects candidates from serving soldiers of Indian Army Navy and Air Force through a selection process which undergoes three phases of tests namely Written Phase SSB Interview and Medical Test The first gate of entry i e Written Test consists of four papers revolving around the topics like General Mental Ability Current General Awareness Interactive Communicative English and Academic Contest Test respectively The new edition of Indian Army ACC Entrance Exam is a Complete Study Package for aspirants preparing for the forthcoming exam It covers the Chapterwise Theory of the entire subjects along with more than 3500 Multiple Choice Questions to give a focused conceptual understanding to the aspirants Model Papers have also been given for the self assessment Structured as per latest syllabus and exam pattern it will support aspirants greatly with its ample amount of study material Table of Contents Model Papers Paper I General Mental Ability Test Reasoning Ability Numerical Ability Paper II Current General Awareness India and World Paper III Interactive Communicative English Objective English Descriptive English Paper IV Academic Contest Test Mathematics General Science Physics Chemistry Biology Humanities

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Practical Exercises in Electrochemistry Felix Oettel, 1897

Contemporary Chemical Analysis Judith F. Robinson, Kenneth A. Robinson, 1998 This book provides a readable yet rigorous introduction to analytical methods with a focus on problem solving skills It stresses the fundamental concepts of chemical analysis and through examples from current journals and other science media shows how the principles and practice of analytical chemistry are used to produce answers to questions in all areas of scientific study and practice Features a balance of topics that is closer to contemporary analytical practice than those covered by other books Introduces

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