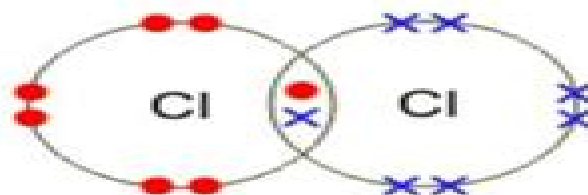


Covalent bonding

A non-metal metal and another non-metal can bond together by **sharing** electrons so that both electrons can achieve a **full outer shell**. This is called covalent bonding. To represent the atoms involved in the bonding we draw crosses for the electrons on one of the atoms and dots for the electrons on the other atoms as shown here.



Instructions

NB: It may be easiest to draw each atom out individually first to work out how many electrons will be shared

1. Draw the circles (outer shells) and overlap them (like a Venn diagram)
2. Write the symbols of the atoms involved in the middle of each circle
3. Draw the **electrons** for one atom as dots (o) and electrons for the other atom as crosses (x).
4. Any shared electrons are drawn in the **overlapping** section.
5. Count up the electrons for EACH atom and check that they each now have a **full outer shell** (2 electrons for the first shell and 8 for the next 2 shells)

TASK: Draw dot and cross diagrams for the following covalent compounds.

F ₂		CH ₄	
HCl		CO ₂	
H ₂ S		H ₂ O	
HF		NH ₃	

Challenge - O₂

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Table of Contents Metallic Bonding Ws Answers

1. Understanding the eBook Metallic Bonding Ws Answers
 - The Rise of Digital Reading Metallic Bonding Ws Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Metallic Bonding Ws Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Metallic Bonding Ws Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Metallic Bonding Ws Answers
 - Personalized Recommendations
 - Metallic Bonding Ws Answers User Reviews and Ratings
 - Metallic Bonding Ws Answers and Bestseller Lists

5. Accessing Metallic Bonding Ws Answers Free and Paid eBooks
 - Metallic Bonding Ws Answers Public Domain eBooks
 - Metallic Bonding Ws Answers eBook Subscription Services
 - Metallic Bonding Ws Answers Budget-Friendly Options
6. Navigating Metallic Bonding Ws Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Metallic Bonding Ws Answers Compatibility with Devices
 - Metallic Bonding Ws Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Metallic Bonding Ws Answers
 - Highlighting and Note-Taking Metallic Bonding Ws Answers
 - Interactive Elements Metallic Bonding Ws Answers
8. Staying Engaged with Metallic Bonding Ws Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Metallic Bonding Ws Answers
9. Balancing eBooks and Physical Books Metallic Bonding Ws Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Metallic Bonding Ws Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Metallic Bonding Ws Answers
 - Setting Reading Goals Metallic Bonding Ws Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Metallic Bonding Ws Answers
 - Fact-Checking eBook Content of Metallic Bonding Ws Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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