

# CHEMISTRY I EQUATION SHEET

## Solubility Rules

Solubility Rules			Form a Compound which is
<b>Negative Ion Plus Positive Ion</b>			
Any neg. ion + group 1a ions Soluble			
Any neg. ion + Ammonium ion Soluble			
Nitrate + any positive ion Soluble			
Acetate	+	any positive ion except Ag <sup>+</sup> or Hg <sup>2+</sup>	Soluble
bromide or iodide + Any other positive ion	+	Ag <sup>+</sup> , Pb <sup>2+</sup> , Hg <sup>2+</sup> or Cu <sup>+</sup>	Not Soluble
Sulfate	+	Ca <sup>2+</sup> , Sr <sup>2+</sup> , Ba <sup>2+</sup> , Ra <sup>2+</sup> , Ag <sup>+</sup> or Pb <sup>2+</sup>	Not Soluble
	+	Any other positive ion	Soluble
Sulfide	+	group 1a ions or ammonium	Soluble
	+	Be <sup>2+</sup> , Mg <sup>2+</sup> , Ca <sup>2+</sup> , Sr <sup>2+</sup> , Ba <sup>2+</sup> or Ra <sup>2+</sup>	Soluble
	+	Any other positive ion	Not Soluble
Hydroxide	+	group 1a ions or ammonium	Soluble
	+	Any other positive ion	Not Soluble
Carbonate, Chromate, Phosphate, or Sulfite + Any other positive ion	+	group 1a ions or ammonium	Soluble
	+	Any other positive ion	Not Soluble

## Equations

D = m/v	c = specific heat
Molarity = $\frac{\text{mol}}{\text{V}}$	d = density
$P_1 V_1 = P_2 V_2$	I = initial amount
$\frac{P_1}{T_1} = \frac{P_2}{T_2}$	F = final amount
Q = mc ΔT	m = mass
nRTPV =	M = molarity
F = I (1/2) *	n = number of moles & number of half lives
n = Total Time	P = pressure
Half Life Time	R = ideal gas constant
	Q = heat energy
	T = temperature
	V = volume
	X = amount of matter left

## Constants

	1 atm = 101.3 kPa Pressure
	1 atm = 760 mm Hg
	R = 8.314 L x kPa / mol°C Universal
Constant or 0.0821 L x atm / mol°C	
Avogadro's number	6.02 X
Specific Heats (J / g°C)	
	Water = 4.180
	Aluminum = 0.895
	Copper = 0.387
	Iron = 0.448

## Activity Series in Decreasing Activity

Metals
Lithium
Rubidium
Potassium
Calcium
Sodium
Magnesium
Aluminum
Manganese
Zinc
Iron
Nickel
Tin
Lead
Hydrogen
Copper
Silver
Platinum
Gold
Nonmetals
Fluorine
Chlorine
Bromine
Iodine

## Polyatomic Ions

Name	Formula
Ammonium	NH <sub>4</sub> <sup>+</sup>
Acetate	C <sub>2</sub> H <sub>3</sub> O <sub>2</sub> <sup>-</sup>
Carbonate	CO <sub>3</sub> <sup>-2</sup>
Bicarbonate or Hydrogen carbonate	HCO <sub>3</sub> <sup>-</sup>
Hypochlorite	ClO <sup>-</sup>
Chlorite	ClO <sub>2</sub> <sup>-</sup>
Chlorate	ClO <sub>3</sub> <sup>-</sup>
Perchlorate	ClO <sub>4</sub> <sup>-</sup>
Hydroxide	OH <sup>-</sup>
Cyanide	CN <sup>-</sup>
Nitrate	NO <sub>3</sub> <sup>-</sup>
Nitrite	NO <sub>2</sub> <sup>-</sup>
Chromate	CrO <sub>4</sub> <sup>-2</sup>
Oxalate	C <sub>2</sub> O <sub>4</sub> <sup>-2</sup>
Sulfate	SO <sub>4</sub> <sup>-2</sup>
Sulfite	SO <sub>3</sub> <sup>-2</sup>
Hydrogen sulfate	HSO <sub>4</sub> <sup>-</sup>
Phosphate	PO <sub>4</sub> <sup>-3</sup>
Phosphite	PO <sub>3</sub> <sup>-3</sup>
Hydrogen phosphate	HPO <sub>4</sub> <sup>-2</sup>

## Metric Conversions

Unit	Abbr.	Meaning
Mega-	M	10 <sup>6</sup>
kilo-	k	10 <sup>3</sup>
hecto-	h	10 <sup>2</sup>
deca-	da	10 <sup>1</sup>
Liter	L	1
gram	g	1
meter	m	1
deci-	d	10 <sup>-1</sup>
centi-	c	10 <sup>-2</sup>
milli-	m	10 <sup>-3</sup>
micro-	μ	10 <sup>-6</sup>
nano-	n	10 <sup>-9</sup>

## Calorie Conversions

Water = 1 cal/ g°C
1000 cal = 1 Cal

# High School Chemistry Formulas Cheat Sheet

**PT Brinkman**

A red circular graphic with a gradient, appearing as a partial circle or a stylized arrow pointing to the right, located to the right of the author's name.

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*Journal of Chemical Education*, 1925 Includes Report of New England Association of Chemistry Teachers and Proceedings of the Pacific Southwest Association of Chemistry Teachers     *A Tentative List of Objectives in the Teaching of Junior High School Mathematics* Raleigh Schorling, 1925     **Chemistry Education** Javier García-Martínez, Elena Serrano-Torregrosa, 2015-02-17 Winner of the CHOICE Outstanding Academic Title 2017 Award This comprehensive collection of top level contributions provides a thorough review of the vibrant field of chemistry education Highly experienced chemistry professors and education experts cover the latest developments in chemistry learning and teaching as well as the pivotal role of chemistry for shaping a more sustainable future Adopting a practice oriented approach the current challenges and opportunities posed by chemistry education are critically discussed highlighting the pitfalls that can occur in teaching chemistry and how to circumvent them The main topics discussed include best practices project based education blended learning and the role of technology including e learning and science visualization Hands on recommendations on how to optimally implement innovative strategies of teaching chemistry at university and high school levels make this book an essential resource for anybody interested in either teaching or learning chemistry more effectively from experience chemistry professors to secondary school teachers from educators with no formal training in didactics to frustrated chemistry students     U Can: Chemistry I For Dummies John T. Moore, Chris Hren, Peter J. Mikulecky, 2015-08-10 Now you can score higher in chemistry Every high school requires a course in chemistry for graduation and many universities require the course for majors in medicine engineering biology and various other sciences U Can Chemistry I For Dummies offers all the how to content you need to enhance your classroom learning simplify complicated topics and deepen your understanding of often intimidating course material Plus you ll find easy to follow examples and hundreds of practice problems as well as access to 1 001 additional Chemistry I practice problems online As more and more students enroll in chemistry courses the need for a trusted and accessible resource to aid in study has never been greater That s where U Can Chemistry I For Dummies comes in If you re struggling in the classroom this hands on friendly guide makes it easy to conquer chemistry Simplifies basic chemistry principles Clearly explains the concepts of matter and energy atoms and molecules and acids and bases Helps you tackle problems you may face in your Chemistry I course Combines how to with try it to form one perfect resource for chemistry students If you re confused by chemistry and want to increase your chances of scoring your very best at exam time U Can Chemistry I For Dummies shows you that you can     Chemical Symbolism and Calculations Stanley Wallace Morse, 1953     **Standard Catalog for High School Libraries**, 1928 The 1st ed accompanied by a list of Library of Congress card numbers for books except fiction pamphlets etc which are included in the 1st ed and its supplement 1926 29     **Industrial & Engineering Chemistry**, 1923     *Teaching First Year Chemistry* Joseph Otto Frank, 1927     **First Revision of Bibliography of Educational Measurements** Indiana University. School of Education. Division of Research

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*Handbook* Maurice Gibbons,2003-02-03 The Self Directed Learning Handbook offers teachers and principals an innovative  
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designed to nurture their natural passion for learning and mastery challenging them to go beyond the easy and familiar so  
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diverse abilities to design and pursue independent course work special projects or even artistic presentations community  
field work or apprenticeships Using this approach the students take on an increasingly autonomous self directed role as they  
progress The heart of the program is the action contract or learning agreement whereby the student sets challenging yet  
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**Governing the Admission of Candidates Into the United States Naval Academy as Midshipmen** ,1934      **School**  
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