

# Elementary Linear Algebra



**Ron Larson**

**Seventh Edition**

# Elementary Linear Algebra Larson 7th Edition

**SA Dillow**



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web limiting reactant theoretical and percentage yields key  $4\text{KO}_2 + 2\text{H}_2\text{O} \rightarrow 4\text{KOH} + 3\text{O}_2$  g if a reaction vessel contains 0 mol  $\text{KO}_2$  and 0 mol  $\text{H}_2\text{O}$  what is the limiting reactant

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web 1 csun.edu/hcchm001/introchemhandouts.html limiting reagents theoretical actual and percent yields example of a limiting reagent problem how many grams of  $\text{NH}_3$  can be produced theoretically from the reaction of 5.0 g of  $\text{N}$  what is the limiting reagent if 8.52 g are actually formed what is the

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web percent yield is a measure of the efficiency of the experimental design yield efficiency  $\frac{\text{mass of product obtained}}{\text{calculated mass of the product expected}} \times 100$  in this reaction a yield 80 is anticipated

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web a write a balanced equation for the reaction  $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$  4  $\text{Al} + 3\text{O}_2 \rightarrow 2\text{Al}_2\text{O}_3$  b determine the theoretical yield of  $\text{Al}_2\text{O}_3$  160.0 g  $\text{Al} \times \frac{1 \text{ mol Al}}{27 \text{ g Al}} \times \frac{2 \text{ mol Al}_2\text{O}_3}{4 \text{ mol Al}} \times 101.96 \text{ g Al}_2\text{O}_3 = 302.3 \text{ g Al}_2\text{O}_3$  c determine the percent yield  $\frac{\text{yield actual} \times 100}{260.0 \text{ g}} \times \frac{100}{86.01} = \frac{\text{theoretical}}{302.3 \text{ g}}$

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web 7  $\text{Na}_2\text{CO}_3 + 4\text{HCl} \rightarrow 2\text{NaCl} + \text{CO}_2 + 2\text{H}_2\text{O}$  is the limiting reactant percent yield 86.6% only four molecules can be made 9 this amount cannot be weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12  $\text{C}$  is the limiting reagent 4.33 g of  $\text{H}_2$  are left over 13

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web the percent yield is calculated as follows  $\text{percent yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100$  percent yield  $\frac{\text{theoretical yield}}{\text{actual yield}} \times 100$  based on this definition we would expect a percent yield

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