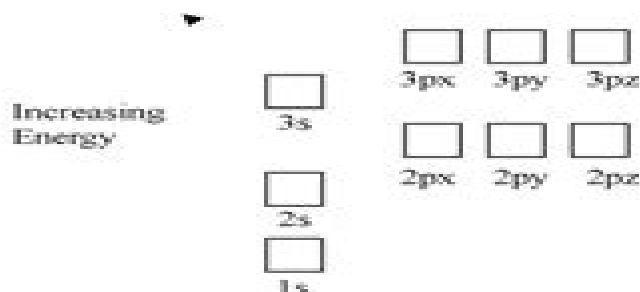


The area in which there is 90% probability of finding an electron is called an orbital. An orbital can hold two electrons.

These orbitals can be represented in the following way:

Appendix B



Each box represents an orbital. Electrons are designated by arrows pointing up or down which represent clockwise or counterclockwise spins.



To determine where the electrons of a particular element are located you use the following rules:

RULE 1: An electron will always be found in the lowest energy orbital available that has room.

Therefore the electron configuration of Hydrogen is:  This is designated as: $1s^1$

RULE 2: An orbital may hold a maximum of 2 electrons.

RULE 3: The Pauli Exclusion Principle – Two electrons sharing an orbital will have opposite spins.

Therefore the electron configuration of Helium is:  This is designated as: $1s^2$

1. Draw an orbital diagram and write the electron configuration for Lithium, Beryllium, and Boron.



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Table of Contents Electron Configuration Packet Key

1. Understanding the eBook Electron Configuration Packet Key
 - The Rise of Digital Reading Electron Configuration Packet Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Electron Configuration Packet Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Electron Configuration Packet Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Electron Configuration Packet Key
 - Personalized Recommendations

- Electron Configuration Packet Key User Reviews and Ratings
- Electron Configuration Packet Key and Bestseller Lists
- 5. Accessing Electron Configuration Packet Key Free and Paid eBooks
 - Electron Configuration Packet Key Public Domain eBooks
 - Electron Configuration Packet Key eBook Subscription Services
 - Electron Configuration Packet Key Budget-Friendly Options
- 6. Navigating Electron Configuration Packet Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Electron Configuration Packet Key Compatibility with Devices
 - Electron Configuration Packet Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Electron Configuration Packet Key
 - Highlighting and Note-Taking Electron Configuration Packet Key
 - Interactive Elements Electron Configuration Packet Key
- 8. Staying Engaged with Electron Configuration Packet Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Electron Configuration Packet Key
- 9. Balancing eBooks and Physical Books Electron Configuration Packet Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Electron Configuration Packet Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Electron Configuration Packet Key
 - Setting Reading Goals Electron Configuration Packet Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Electron Configuration Packet Key
 - Fact-Checking eBook Content of Electron Configuration Packet Key

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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