

5

ELECTRONS IN ATOMS

SECTION 5.1 MODELS OF THE ATOM (pages 127–132)

This section summarizes the development of atomic theory. It also explains the significance of quantized energies of electrons as they relate to the quantum mechanical model of the atom.

► The Development of Atomic Models (pages 127–128)

1. Complete the table about atomic models and the scientists who developed them.

Scientist	Model of Atom
Dalton	The atom is a solid indivisible mass.
Thomson	The atom is a ball of positive charge with electrons stuck into the ball.
Rutherford	Most of an atom's mass is concentrated in the small, positively charged nucleus. The electrons surround the nucleus and the rest of the atom is empty space.
Bohr	Electrons are arranged in concentric circular paths around the nucleus.

2. Is the following sentence true or false? The electrons in an atom can exist between energy levels. false

► The Bohr Model (pages 128–129)

3. What is a small, discrete unit of energy called?
It is called a quantum.
4. Circle the letter of the term that completes the sentence correctly: A quantum of energy is the amount of energy required to
- move an electron from its present energy level to the next lower one
 - maintain an electron in its present energy level
 - c.** move an electron from its present energy level to the next higher one
5. In general, the higher the electron is on the energy ladder, the farther it is from the nucleus.

Electron Arrangement In Atoms Guided Reading Answers

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Electron Arrangement In Atoms Guided Reading Answers:

Prentice Hall Chemistry, 2000 **AQA AS Chemistry: Unit 1 Foundation Chemistry Student Unit Guide Ebook**

Margaret Cross, 2008-09-26 Student Unit Guides are perfect for revision Each guide is written by an examiner and explains the unit requirements summarises the relevant unit content and includes a series of specimen questions and answers There are three sections to each guide Introduction includes advice on how to use the guide an explanation of the skills being tested by the assessment objectives an outline of the unit or module and depending on the unit suggestions for how to revise effectively and prepare for the examination questions Content Guidance provides an examiner s overview of the module s key terms and concepts and identifies opportunities to exhibit the skills required by the unit It is designed to help students to structure their revision and make them aware of the concepts they need to understand the exam and how they might analyse and evaluate topics Question and Answers sample questions and with graded answers which have been carefully written to reflect the style of the unit All responses are accompanied by commentaries which highlight their respective strengths and weaknesses giving students an insight into the mind of the examiner **Addison-Wesley Chemistry** Antony C.

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Holt Physical Science Mapi M. Cuevas, 1994

The Pearson Guide To Objective Physics For The Iit-Jee, 2/E Dudeja, 2009-09

Prentice Hall Physical Science Concepts in Action Program Planner

National Chemistry Physics Earth Science, 2003-11 Prentice Hall Physical Science Concepts in Action helps students make the important connection between the science they read and what they experience every day Relevant content lively explorations and a wealth of hands on activities take students understanding of science beyond the page and into the world around them Now includes even more technology tools and activities to support differentiated instruction

Matter Prentice-Hall Staff, 1994

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by David E. Goldberg Margaret G. Kimble,1994 *Student's Guide to Fundamentals of Chemistry* Jo A. Beran,2013-10-22 Student s Guide to Fundamentals of Chemistry Fourth Edition provides an introduction to the basic chemical principles This book deals with various approaches to chemical principles and problem solving in chemistry Organized into 25 chapters this edition begins with an overview of how to define and recognize the more common names and symbols in chemistry This text then discusses the historical development of the concept of atom as well as the historical determination of atomic weights for the elements Other chapters consider how to calculate the molecular weight of a compound from its formula This book discusses as well the characteristics of a photon in terms of its particle like properties and defines the wavelength frequency and speed of light The final chapter deals with the fundamental components of air and the classification of materials formed in natural waters This book is a valuable resource for chemistry students lecturers and instructors **Chemistry of Matter**

,1993 As the reader read s this textbook he will learn about the interactions of matter that can occur in a test tube in nature and even inside himself P 9 **Chemistry: Concepts and Problems** Clifford C. Houk,Richard Post,1996-03-09

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