

Before beginning this experiment in the laboratory, you should be able to answer the following questions.

1. Define the following terms: *faraday*, *salt bridge*, *anode*, *cathode*, *voltic cell*, and *electrolytic cell*.

2. Write a chemical equation for the reaction that occurs in the following cell:  $\text{Cu} | \text{Cu}^{2+}(\text{aq}) || \text{Ag}^{+}(\text{aq}) | \text{Ag}$ .

3. Given the following  $E^{\circ}$ 's, calculate the standard-cell potential for the cell in question 2.



4. Calculate the voltage of the following cell:



5. Calculate the cell potential, the equilibrium constant, and the free-energy change for



given the following  $E^{\circ}$  values:



# Electrochemical Lab Pre Lab Answers

**Kevin Karplus**



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