

10.3

THE NATURE OF SOLIDS

SECTION REVIEW

Objectives

- Describe how the degree of organization of particles distinguishes solids from gases and liquids
- Distinguish between a crystal lattice and a unit cell
- Explain how allotropes of an element differ

Key Terms

- melting point
- crystal
- unit cell
- allotropes
- amorphous solids
- glasses

Part A Completion

Use your completion material to check your understanding of the concepts and terms discussed in this section. Each blank can be completed with a name, short phrase, or number.

Solids tend to be dense and _____. They do not flow or take the shape of their containers, like liquids do, because the particles in solids vibrate around _____ points. When a solid is sufficiently heated and its particles vibrate so strongly that they are no longer held in fixed positions, the solid _____. The _____ is the temperature at which a solid turns to liquid. The melting and _____ of a substance are at the same temperature. In general, ionic solids tend to have relatively _____ melting points while molecular solids tend to have relatively low melting points. Most solids are _____ in nature. The particles are arranged in a pattern known as a crystal _____. The smallest unit of a crystal lattice is the _____. Some solids have a disordered array of particles and are _____.

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Nature Of Solids Section Review Key

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