IB SL and HL Topic 1 to 5 Chemistry MCQ 40 Marks

- What amount of oxygen, O₂, (in moles) contains 1.8×10²² molecules?
 - A. 0.0030
 - B. 0.030
 - C. 0.30
 - D. 3.0

2. Which compound has the empirical formula with the greatest mass?

- A. C₂H₆.
- B. C₄H₁₀
- C. C_5H_{10}
- D. C₆H₆

(Total I mark)

(Total I mark)

3. $C_2H_2(g) + O_2(g) \rightarrow CO_2(g) + H_2O(g)$

When the equation above is balanced, what is the coefficient for oxygen?

- Λ. 2
- B. 3
- C. 4
- D. 5

(Total I mark)

3.0 dm³ of sulfur dioxide is reacted with 2.0 dm³ of oxygen according to the equation below.

$$2SO_2(g) + O_2(g) \rightarrow 2SO_3(g)$$

What volume of sulfur trioxide (in dm³) is formed? (Assume the reaction goes to completion and all gases are measured at the same temperature and pressure.)

- A. 5.0
- B. 4.0
- C. 3.0
- D. 2.0

(Total I mark)

<u>Ib Chemistry Hl 2013 Paper 1 Per</u>

Alexandra Juniper, Jason Murgatroyd

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