

1. The following reaction of acetylene is used in a single-headed torch:

2. Change in ΔH°
 3. ΔG°



4. ΔH° of formation of CO_2 is -393.5 kJ/mol .
 5. ΔH° of formation of H_2O is -285.8 kJ/mol .
 6. ΔH° of formation of C_2H_2 is $+226.7 \text{ kJ/mol}$.

7. Calculate the ΔH° of the reaction of acetylene with oxygen.

8. Which side of the torch is hotter? Is it used for welding or cutting?

9. If you, a student, had to choose a single-headed torch for welding, which one would you choose? Why? (Explain your choice.)

10. The reaction of acetylene with oxygen is exothermic.
 11. The reaction of acetylene with oxygen is endothermic.
 12. The reaction of acetylene with oxygen is exothermic.
 13. The reaction of acetylene with oxygen is endothermic.

14. The reaction of acetylene with oxygen is exothermic.
 15. The reaction of acetylene with oxygen is endothermic.
 16. The reaction of acetylene with oxygen is exothermic.
 17. The reaction of acetylene with oxygen is endothermic.

18. If you, a student, had to choose a single-headed torch for welding, which one would you choose? Why? (Explain your choice.)



19. ΔH° of formation of CO_2 is -393.5 kJ/mol .
 20. ΔH° of formation of H_2O is -285.8 kJ/mol .
 21. ΔH° of formation of C_2H_2 is $+226.7 \text{ kJ/mol}$.

22. If you, a student, had to choose a single-headed torch for welding, which one would you choose? Why? (Explain your choice.)



23. ΔH° of formation of CO_2 is -393.5 kJ/mol .
 24. ΔH° of formation of H_2O is -285.8 kJ/mol .
 25. ΔH° of formation of C_2H_2 is $+226.7 \text{ kJ/mol}$.

Genetics Problem Set 3 Answers

M Woodhall



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