		Ю		н	
		н	ш	ш	

SS NUMBER - total amount

An atom is a really tiny particle. It has a radius of about 0.1 nanometres

(1x10-10 m) Atoms are neutral because the no. protons

= no. electrons An atom consists of:

Particle	Relative	Charg
Proton	1	+1
Neutron	1	0
Electron	0 (very	-1

You can find the

of Mumber of mestrons = atomic

THE NUCLEUS

- •Middle of the atom
- Contains protons and neutrons
 - Has a radius of around 1x10-14m (1/100° of atom radius)
- Positively charged due to protons
- Almost whole mass of atom concentrated in mucleus.

THE ELECTRONS

Move around nucleus in electron shells

- •Negatively charged & tiny, but cover a lot of space
- •The volume of their orbits determines the size
- of the atoms Negatively charged
- ·Bond to other electrons to make chemical bond



Element

A substance made up of atoms that all have the same number of protons in their nucleus (i.e. all are the same atom)

Isotopes

Different forms of the same element, which have the same number of protons but a different number of neutrons.

2 common

examp -Carbon-12

*Carbon-13

Because many elements can exist as a

number of

different isotopes... relative atomic

mass (A) is used instead of mass number

when referring

to the element

Carbon-13 Carlotte. and Property. ALL PROPERTY. # s. Streetman in Thiredness Machine anadom Markett breaker 18.16 44.47

+ 111

relative alonic mass =

sum of all isotope abundances

- 10

To balance an equation, you should write out all the elements in the middle, write out how many there are on either side and add numbers to balance it

CHEMISTRY CT

Atomic structure, separating

H2SO4 + NaOH Na/50++ H₂O

BQ) Copper has two stable isotopes. Cu-63 has an abundance of 69.2% and Cu-65 has an abundance of 30.8%. Calculate the relative atomic mass of copper to 1dp.

(882-1)-(383-65)_ 61584-2981_43614 63.616 = 63.6

to enable

num off instage abundance + instage mumber's

The relative atomic mass is an average mass taking into account the different masses & abundances (amounts) of all the isotopes that make up the

Compound

Substances formed from 2/more elements, where the atoms of each are held together by chemical bonds.

- Making bond involves atoms giving away, taking or sharing electrons
- A chemical bond is required to break these bonds and separate the compound
- There are 2 types of compound bonds:

1) IONIC BONDING

- · When a metal and a non-metal react to form compound
- The metal loses electrons to form a positive ion
- The non-metal gains electrons to form a negative ion
- The opposite charges of the ions. means they're strongly attracted

2) COVALENT BONDING

- A compound formed from nonmetal consists of molecules
- Each atoms shares electrons with another atom

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